

SEARCH REQUEST FORM

CIN 1/30/02

Examiner # (Mandatory): Cam Nguyen Requester's Full Name: 58778Art Unit 1754 Location (Bldg/Room#): CP3 - 9B06 Phone (circle) 205 306 308 3923Serial Number: 09/653,719 Results Format Preferred (circle): PAPER DISK E-MAILTitle of Invention Fischer-Tropsch catalyst enhancementInventors (please provide full names): Alla Jurjena, Krylova et al.Earliest Priority Date: 9/1/00

Keywords (include any known synonyms, registry numbers, explanation of initialisms):

*Please search for claims 1-17 only!

see copy of claims attached.

Search Topic:

Please write detailed statement of the search topic, and the concept of the invention. Describe as specifically as possible the subject matter to be searched. Define any terms that may have a special meaning. Give examples of relevant citations, authors, etc., if known. You may include a copy of the abstract and the broadcast or most relevant claim(s).

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L7 6691 SEA FILE=HCAPLUS ABB=ON DAM OR DISPERS?(4A)ACTIV?(2A)METAL?
L9 287 SEA FILE=HCAPLUS ABB=ON L7 AND (CAT/RL OR CATALYST?)
L10 150 SEA FILE=HCAPLUS ABB=ON L9 AND (RE OR RHENIUM OR RU OR
RUTHEN? OR PD OR PALLAD? OR FE OR IRON OR FERRIC OR FERROUS OR
CO OR COBALT)
L11 42 SEA FILE=HCAPLUS ABB=ON L10 AND (HYDROGEN OR H2)
L12 19 SEA FILE=HCAPLUS ABB=ON L11 AND HYDROGENAT?
L13 14 SEA FILE=HCAPLUS ABB=ON L11 AND (PREP OR IMF OR SPN)/RL
L14 1 SEA FILE=HCAPLUS ABB=ON L11 AND FISCHER?
L15 29 SEA FILE=HCAPLUS ABB=ON L11 AND CATALYST?/TI
L16 27 SEA FILE=HCAPLUS ABB=ON L11 AND INORGANIC/SC,SX
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09/653719

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=> d que 125

L19 8537 SEA FILE=WPIX ABB=ON DAM OR DISPERS? (4A)ACTIV? (2A)METAL?
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OR PD OR PALLAD? OR FE OR IRON OR FERRIC OR FERROUS OR CO OR
COBALT)
L21 40 SEA FILE=WPIX ABB=ON L20 AND CATALYST?
L22 13 SEA FILE=WPIX ABB=ON L21 AND (H2 OR HYDROGEN)
L24 5 SEA FILE=WPIX ABB=ON L22 AND (HYDROGENAT? OR FISCHER?)
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PROCESSING COMPLETED FOR L25
L26 52 DUP REM L17 L25 (0 DUPLICATES REMOVED)

=> d all 126 (1-52)

L26 ANSWER 1 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 2001:868305 HCAPLUS
DN 136:8990
TI Olivine-supported nickel catalyst for fluidized-bed gasification
of biomass for manufacture of fuel gases and synthesis gas
IN Courson, Claire; Petit, Corinne; Kienemann, Alain; Foscolo, Pier Ugo;
Rapagna, Sergio; Matera, Domenico Antonio
PA Universite Louis Pasteur, Fr.; ENEA - Ente per le Nuove Tecnologie,
l'Energia e l'Ambiente; Universita degli Studi di l'Aquila
SO PCT Int. Appl., 18 pp.
CODEN: PIXXD2
DT Patent
LA French
IC ICM B01J023-755
ICS C10J003-54; C01B003-38
CC 52-1 (Electrochemical, Radiational, and Thermal Energy Technology)
Section cross-reference(s): 51
FAN.CNT 1

PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
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PI WO 2001089687 A1 20011129 WO 2001-FR1547 20010518
W: AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN,
CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH,
GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR,
LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, PL, PT,
RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, TZ, UA, UG, US,
UZ, VN, YU, ZA, ZW, AM, AZ, BY, KG, KZ, MD, RU, TJ, TM
RW: GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZW, AT, BE, CH, CY,
DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE, TR, BF,
BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG
FR 2809030 A1 20011123 FR 2000-6526 20000522

PRAI FR 2000-6526 A 20000522

AB Catalyst for fluidized-bed gasification of biomass, characterized in that it contains, a active metal (typically Ni) phase dispersed on an olivine support of general formula $(\text{Mg}_x\text{Fe}_{1-x})_2\text{SiO}_4$ ($x = 0-1$), esp. $x = 0.92$, with stoichiometric excess Fe (as oxide). The catalyst is prepd. by: (1) washing the olivine support, with drying and a first calcination, (2) grinding and sieving the support, (3) depositing NiO on the support by impregnation (from one or more Ni salts, such as $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$), (4) a second calcination step, at 900-1300.degree. (preferably .apprx.1100.degree.), (5) sieving the catalyst, and (6) reducing the nickel catalyst (at 750-900.degree., preferably .apprx.865.degree.) in hydrogen to deposit metallic Ni. The catalyst can be used for fluidized-bed carbonization/gasification of biomass, esp. for hydrogen manuf. and synthesis gas reactions by methane reforming (with CO₂ or water) and steam reforming of tar byproduct.

ST fluidized bed gasification biomass nickel catalysis; synthesis gas biomass fluidized bed gasification; fuel gas biomass fluidized bed gasification; olivine support nickel gasification biomass

IT Tar
RL: PNU (Preparation, unclassified); RCT (Reactant); PREP (Preparation); RACT (Reactant or reagent)
(biomass, in-situ formation and reforming of; olivine-supported nickel catalyst for fluidized-bed gasification of biomass for manuf. of fuel gases and synthesis gas)

IT Fuel gas manufacturing
(gasification, fluidized-bed, of biomass; olivine-supported nickel catalyst for fluidized-bed gasification of biomass for manuf. of fuel gases and synthesis gas)

IT Synthesis gas manufacturing
(gasification-steam reforming; olivine-supported nickel catalyst for fluidized-bed gasification of biomass for manuf. of fuel gases and synthesis gas)

IT Reforming
Steam reforming
(in-situ, of biomass tars; olivine-supported nickel catalyst for fluidized-bed gasification of biomass for manuf. of fuel gases and synthesis gas)

IT Reforming catalysts
Steam reforming catalysts
(nickel, for biomass tars; olivine-supported nickel catalyst for fluidized-bed gasification of biomass for manuf. of fuel gases and synthesis gas)

IT Biomass
Fluidized beds
Thermal decomposition
(olivine-supported nickel catalyst for fluidized-bed gasification of biomass for manuf. of fuel gases and synthesis gas)

IT Olivine-group minerals
RL: CAT (Catalyst use); USES (Uses)
(support; olivine-supported nickel catalyst for fluidized-bed
gasification of biomass for manuf. of fuel gases and synthesis gas)

IT Biomass
RL: PNU (Preparation, unclassified); RCT (Reactant); PREP
(Preparation); RACT (Reactant or reagent)
(tar, in-situ formation and reforming of; olivine-supported nickel
catalyst for fluidized-bed gasification of biomass for manuf.
of fuel gases and synthesis gas)

IT 1313-99-1, Nickel oxide (NiO), uses
RL: CAT (Catalyst use); USES (Uses)
(catalyst; olivine-supported nickel catalyst for
fluidized-bed gasification of biomass for manuf. of fuel gases and
synthesis gas)

IT 74-82-8P, Methane, preparation
RL: PNU (Preparation, unclassified); RCT (Reactant); PREP
(Preparation); RACT (Reactant or reagent)
(in-situ formation and reforming of; olivine-supported nickel
catalyst for fluidized-bed gasification of biomass for manuf.
of fuel gases and synthesis gas)

IT 1333-74-0P, Hydrogen, preparation
RL: IMF (Industrial manufacture); PREP (Preparation)
(manuf. of; olivine-supported nickel catalyst for
fluidized-bed gasification of biomass for manuf. of fuel gases and
synthesis gas)

IT 13478-00-7, Nickel nitrate, hexahydrate
RL: RCT (Reactant); RACT (Reactant or reagent)
(nickel source; olivine-supported nickel catalyst for
fluidized-bed gasification of biomass for manuf. of fuel gases and
synthesis gas)

IT 109166-61-2, Iron magnesium silicate ((Fe,Mg)2(SiO4))
RL: CAT (Catalyst use); USES (Uses)
(olivine, support; olivine-supported nickel catalyst for
fluidized-bed gasification of biomass for manuf. of fuel gases and
synthesis gas)

RE.CNT 1 THERE ARE 1 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

(1) Basf Ag; FR 2121148 A 1972 HCAPLUS

L26 ANSWER 2 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 2001:936972 HCAPLUS
TI Hydrodechlorination of 1,1-dichlorotetrafluoroethane on supported
palladium catalysts. A static-circulation reactor study
AU Karpinski, Zbigniew; D'Itri, Julie L.
CS Department of Catalysis on Metals, Institute of Physical Chemistry, Polish
Academy of Sciences, Warsaw, PL-01224, Pol.
SO Catalysis Letters (2001), 77(1-3), 135-140
CODEN: CALEER; ISSN: 1011-372X
PB Kluwer Academic/Plenum Publishers
DT Journal
LA English
CC 45 (Industrial Organic Chemicals, Leather, Fats, and Waxes)
AB Supported palladium catalysts were studied in CF₃CFC₁₂
hydrodechlorination at 100.degree.C using a static-circulation system. In
order to minimize catalyst's deactivation a large excess of
hydrogen was employed (H₂/CF₃CFC₁₂ ratio 54/1). In
spite of this precaution significant inhibition of the process occurred,
assocd. with blocking palladium surface by hydrogen
chloride species. Differences in the catalytic behavior of

alumina-supported and unsupported palladium are discussed. A mild dependence between the catalytic activity and Pd dispersion was found. The Pd/Al₂O₃ catalyst characterized by low metal dispersion was more active than highly dispersed catalysts, showing the overall activity and selectivity to CF₃CFH₂ comparable with those obsd. by other authors for palladium single crystals. It is speculated that the most active sites for hydrodechlorination are plane atoms, whereas low coordination sites (on edges and corners of metal crystallites) are less suitable.

RE.CNT 21 THERE ARE 21 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

- (1) Basset, J; J Catal 1970, V16, P53 HCPLUS
- (2) Campbell, J; Trans Faraday Soc 1963, V59, P2583 HCPLUS
- (3) Coq, B; J Catal 1986, V101, P434 HCPLUS
- (4) Coq, B; J Catal 1993, V141, P21 HCPLUS
- (5) Coq, B; J Phys Chem 1995, V99, P11159 HCPLUS
- (6) Early, K; J Catal 1999, V182, P219 HCPLUS
- (7) Gervasutti, C; J Fluorine Chem 1981, V19, P1 HCPLUS
- (8) Ichikawa, S; J Catal 1985, V91, P1 HCPLUS
- (9) Juszczuk, W; Appl Catal A 1998, V166, P311 HCPLUS
- (10) Karpinski, Z; J Catal 1996, V164, P378 HCPLUS
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- (13) Rupprechter, G; Catal Lett 1997, V48, P17 HCPLUS
- (14) Tanaka, M; J Catal 1970, V16, P157 HCPLUS
- (15) Thomson, J; J Mol Catal 1991, V67, P117 HCPLUS
- (16) Thomson, J; J Mol Catal 1991, V68, P347 HCPLUS
- (17) van Hardeveld, R; Surf Sci 1969, V15, P189 HCPLUS
- (18) van de Sandt, E; Appl Catal A 1997, V155, P59 HCPLUS
- (19) van de Sandt, E; Catal Today 1997, V35, P163 HCPLUS
- (20) van de Sandt, E; Recl Trav Chim Pays-Bas 1996, V115, P505 HCPLUS
- (21) Wiersma, A; Catal Today 1996, V27, P257 HCPLUS

L26 ANSWER 3 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 2001:499154 HCPLUS

DN 135:359907

TI Hydroisomerization of n-butane over hybrid catalysts

AU Canizares, P.; Dorado, F.; Sanchez-Herrera, P.

CS Faculty of Chemistry, Department of Chemical Engineering, University of Castilla-La Mancha, Ciudad Real, 13004, Spain

SO Appl. Catal., A (2001), 217(1-2), 69-78

CODEN: ACAGE4; ISSN: 0926-860X

PB Elsevier Science B.V.

DT Journal

LA English

CC 51-6 (Fossil Fuels, Derivatives, and Related Products)

AB The influence of redn. temp., metal incorporation technique, and metal loading on the hydroisomerization of n-butane over hybrid catalysts prep'd. by phys. mixing H-mordenite (HMOR) and Pd /montmorillonite (Mont) was studied. Temp.-programmed desorption of ammonia (TPDA), Fourier transform IR (FTIR), at. absorption spectroscopy (AA), BET surface area/pore size distributions, and chemisorption measurements were used to characterize the catalysts. The method used to load the metal was found to be an important factor that affected the final catalyst acid properties, metal dispersion and, hence, its activity and selectivity. Thus, the highest selectivity to isobutane was achieved over an ion-exchanged catalyst. Compared to classic metal-supported HMOR catalysts, hybrid catalysts showed a higher selectivity to isobutane.

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ST butane hydroisomerization **catalyst**
IT Hydrogen mordenite-type zeolites
RL: CAT (Catalyst use); USES (Uses)
(contg. Pd and montmorillonite; hydroisomerization of
n-butane over hybrid **catalysts**)
IT Petroleum hydrotreating **catalysts**
(hydroisomerization; hydroisomerization of n-butane over hybrid
catalysts)
IT 1318-93-0, Montmorillonite, uses
RL: CAT (Catalyst use); USES (Uses)
(contg. Pd and H-mordenite; hydroisomerization of n-butane
over hybrid **catalysts**)
IT 7440-05-3, Palladium, uses
RL: CAT (Catalyst use); USES (Uses)
(contg. montmorillonite and H-mordenite; hydroisomerization of n-butane
over hybrid **catalysts**)
IT 106-97-8, Butane, reactions
RL: PEP (Physical, engineering or chemical process); RCT (Reactant); PROC
(Process)
(hydroisomerization of n-butane over hybrid **catalysts**)

RE.CNT 26 THERE ARE 26 CITED REFERENCES AVAILABLE FOR THIS RECORD

RE

- (1) Asuquo, R; J Catal 1995, V155, P376 HCPLUS
- (2) Beyer, H; Zeolites 1998, V8, P79
- (3) Canizares, P; Appl Catal A: Gen 1998, V169, P137 HCPLUS
- (4) Canizares, P; Appl Catal A: Gen 2000, V196, P225 HCPLUS
- (5) Canizares, P; Appl Catal A: Gen 2000, V190, P233 HCPLUS
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- (10) Guisnet, M; Appl Catal A: Gen 1992, V87, P255 HCPLUS
- (11) Jia, C; Appl Catal A: Gen 1998, V106, P185
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- (13) Krannila, H; J Catal 1992, V135, P115 HCPLUS
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- (15) Liu, H; Appl Catal A: Gen 1996, V137, P167 HCPLUS
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- (18) Narayanan, S; Appl Catal A: Gen 1998, V167, P103 HCPLUS
- (19) Parton, R; Appl Catal A: Gen 1991, V76, P131 HCPLUS
- (20) Rodriguez, F; J Catal 1986, V99, P171
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- (25) Taylor, R; Appl Catal A: Gen 1994, V119, P121 HCPLUS
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L26 ANSWER 4 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 2001:539072 HCPLUS

DN 135:244717

TI Environmental **catalysts** containing Ni and Mo supported on
zeolite

AU Peng, Ming-sheng; Li, Di-en

CS Inst. Gemstones Mineral Materials, Zhongshan Univ., Canton, 510275, Peop.
Rep. China

SO Anquan Yu Huanjing Xuebao (2001), 1(2), 42-44

CODEN: AYHXA2; ISSN: 1009-6094

PB Anquan Yu Huanjing Xuebao Zazhishe

DT Journal
LA Chinese
CC 51-9 (Fossil Fuels, Derivatives, and Related Products)
AB Catalytic **hydrogenation** and hydrocracking (HC) are refining processes for petroleum and coal liqs. The most popular **catalysts** used in industry are Co-/Ni-promoted Mo/W sulfides supported on .gamma.-Al₂O₃. New **catalysts** are in great demand since the contents of pollutants such as S, N, and heavy metals contained in oil must be reduced to one tenth of the present level. For such a purpose, the effects of B on the structure and activity of Ni-Mo/.gamma.-Al₂O₃ **catalysts** are studied in the present paper. .gamma.-Al₂O₃-B₂O₃ supports and Ni-Mo/.gamma.-Al₂O₃-B₂O₃ **catalysts** in which the contents of NiO and MoO₃ are 3 and 12%, resp., are prep'd. by impregnation method. Di-Ph and cyclohexylbenzene are formed in the hydrodesulfurization (HDS) of dibenzothiophene (DBT). X-ray diffraction and X-ray photo-electron spectrum are used to analyze the surface acidity and dispersion of metals. It is shown that the addn. of B modifies the acidity and the **active metal dispersion** of the **catalysts** and thus enhances the catalytic hydroprocessing activities of these model compds. over the Ni-Mo/.gamma.-Al₂O₃ **catalysts**. Activity was at max. for 1% B₂O₃ content. Ultrastable Y-type (USY) zeolite supported Ni-Mo **catalysts** are also studied. NH₄Y zeolite is prep'd. from NaY zeolite by ion exchange with (NH₄)₂SO₄. USY zeolite is obtained after the calcination of NH₄Y zeolite, dealumination, and removing Al by H₂SO₄. **Catalysts** supported on USY zeolite are prep'd. by ion exchange method. These **catalysts** are treated with H₂S-H₂ at 400 .degree.C and used in HDS of DBT and HC of Decalin, diphenylmethane, and Tetralin. USY zeolite-supported Ni-Mo **catalyst** has significantly higher activity.
ST hydrodesulfurization hydrocracking **catalyst**
IT Hydrocracking **catalysts**
Hydrodesulfurization **catalysts**
(Ni-Mo-zeolite)
IT Petroleum hydrotreating **catalysts**
(hydrodesulfurization and hydrocracking of model compds. on Ni-Mo-zeolite)
IT Ultrastable Y zeolites
RL: CAT (**Catalyst use**); USES (**Uses**)
(hydrodesulfurization and hydrocracking of model compds. on Ni-Mo-zeolite **catalyst**)
IT 7439-98-7, Molybdenum, uses 7440-02-0, Nickel, uses
RL: CAT (**Catalyst use**); USES (**Uses**)
(hydrodesulfurization and hydrocracking of model compds. on Ni-Mo-zeolite **catalyst**)
IT 91-17-8, Decalin 101-81-5, Diphenylmethane 119-64-2, Tetralin
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(hydrodesulfurization and hydrocracking of model compds. on Ni-Mo-zeolite **catalyst**)
IT 132-65-0, Dibenzothiophene
RL: REM (Removal or disposal); PROC (Process)
(hydrodesulfurization and hydrocracking of model compds. on Ni-Mo-zeolite **catalyst**)

L26 ANSWER 5 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 2000-246502 [21] WPIX
CR 2000-246501 [19]; 2000-246503 [19]; 2001-353964 [36]
DNC C2000-074580
TI Production of hydrocarbons in Fischer-Tropsch process uses catalyst comprising iron, cobalt, nickel and or ruthenium dispersed in stabilized aquasol of colloidal oxide

of cerium, zirconium, titanium, aluminum and or silicon.

DC E17 H04

IN KOURTAKIS, K; MANZER, L E

PA (CONO) CONOCO INC

CYC 86

PI WO 2000010704 A1 20000302 (200021)* EN 23p B01J021-08

RW: AT BE CH CY DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL
OA PT SD SE SL SZ UG ZW
W: AE AL AM AT AU AZ BA BB BG BR BY CA CH CN CU CZ DE DK EE ES FI GB
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LV MD MG MK MN MW MX NO NZ PL PT RO RU SD SE SG SI SK SL TJ TM TR
TT UA UG UZ VN YU ZA ZW

AU 9955721 A 20000314 (200031)

B01J021-08

EP 1128905 A1 20010905 (200151) EN B01J021-08

R: AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU MC NL PT SE

ADT WO 2000010704 A1 WO 1999-US18895 19990819; AU 9955721 A AU 1999-55721
19990819; EP 1128905 A1 EP 1999-942315 19990819, WO 1999-US18895 19990819

FDT AU 9955721 A Based on WO 2000010704; EP 1128905 A1 Based on WO 2000010704

PRAI US 1999-377008 19990818; US 1998-97192P 19980820; US 1998-97193P
19980820; US 1998-97194P 19980820; US 1999-376873 19990818; US
1999-377007 19990818

IC ICM B01J021-08

ICS C07C027-00

AB WO 2000010704 A UPAB: 20010910

NOVELTY - A catalyst comprising iron, cobalt,
, nickel and or ruthenium as catalytically active
metal dispersed in stabilized aquasol of colloidal oxide
of cerium, zirconium, titanium, aluminum and/or silicon provides high
selectivity to 5+C hydrocarbons and enhances process economics.

DETAILED DESCRIPTION - Production of hydrocarbons involves contacting
a feedstream of hydrogen and carbon monoxide with
catalyst in a reaction zone at conversion conditions. The
catalyst comprises iron, cobalt, nickel and or
ruthenium dispersed in matrix material comprising a derivative of
a stabilized aquasol of colloidal oxide of cerium, zirconium, titanium,
aluminum and/or silicon.

INDEPENDENT CLAIMS are included for:

(a) a Fischer-Tropsch catalyst as above; and

(b) preparing the catalyst by mixing a colloidal Sol of an
oxide of cerium, zirconium, titanium, aluminum and/or silicon with soluble
salt of iron, cerium, nickel and/or ruthenium,
destabilizing the colloid to form a gel and removing solvent from the gel.

USE - For use in Fischer-Tropsch processes.

Dwg.0/0

FS CPI

FA AB; DCN

MC CPI: E10-J02D; E31-P05A; E35-K; E35-L; E35-U; E35-V; E35-W; E35-X;
H04-E05; H04-F02E; N01-C; N01-D02; N02-A; N02-B; N02-C; N02-E01;
N03-A; N03-B

L26 ANSWER 6 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 2000:318441 HCPLUS

DN 133:17152

TI Study on the reaction kinetics of catalytic hydrogenation of
diacetyl monoxime

AU Lu, Yue-qing; Lu, Wei-ming; Wang, Xiang-yu; Zheng, Xiao-ming

CS Institute of Catalysis, Zhejiang University (XiXi Campus), Hangzhou,
310028, Peop. Rep. China

SO Fenzi Cuihua (2000), 14(2), 147-150

CODEN: FECUEN; ISSN: 1001-3555

PB Kexue Chubanshe
DT Journal
LA Chinese
CC 22-7 (Physical Organic Chemistry)
Section cross-reference(s): 67
AB The kinetics of diacetyl monoxime (**DAM**) **hydrogenation** /cyclocondensation to tetramethylpyrazine catalyzed by Pd/PPh₃ has been studied. The reaction follows second order kinetics with respect to **DAM** in the range of investigation, the zeroth order kinetics with respect to H₂ pressure in optimum reaction conditions, and the first order kinetics with respect to **catalyst** concn. The kinetic equation of **hydrogenation** reaction rate obtained is $r=k[\text{DAM}]^2[\text{Cat.}]$. The activation energy for this reaction is E_a = 50.41 kJ .cntdot. mol⁻¹. The inferred mechanism of reaction is in agreement with exptl. results.
ST diacetyl oxime **hydrogenation** cyclocondensation **palladium** phosphine **catalyst** kinetics mechanism
IT Cyclocondensation reaction
Cyclocondensation reaction **catalysts**
Cyclocondensation reaction kinetics
 Hydrogenation
 Hydrogenation catalysts
 Hydrogenation kinetics
 (kinetics and mechanism of Pd/PPh₃-catalyzed **hydrogenation/cyclondensation** of diacetyl monoxime)
IT 603-35-0D, Triphenylphosphine, **palladium** complexes 7440-05-3D,
Palladium, triphenylphosphine complexes
RL: CAT (Catalyst use); USES (Uses)
 (kinetics and mechanism of Pd/PPh₃-catalyzed **hydrogenation/cyclondensation** of diacetyl monoxime)
IT 57-71-6, Diacetyl monoxime
RL: PEP (Physical, engineering or chemical process); PRP (Properties); RCT (Reactant); PROC (Process)
 (kinetics and mechanism of Pd/PPh₃-catalyzed **hydrogenation/cyclondensation** of diacetyl monoxime)
IT 1124-11-4P, Tetramethylpyrazine
RL: SPN (Synthetic preparation); PREP (Preparation)
 (kinetics and mechanism of Pd/PPh₃-catalyzed **hydrogenation/cyclondensation** of diacetyl monoxime)

L26 ANSWER 7 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 2001:60701 HCAPLUS
DN 134:208270
TI Synthesis, characterization and physicochemical properties of polyphenols prepared by cobalt(II) catalysed oxidative polymerization
AU Puzari, A.; Baruah, Jubraj B.
CS Department of Chemistry, Indian Institute of Technology, Guwahati, 781 001, India
SO React. Funct. Polym. (2000), 46(2), 101-107
CODEN: RFPOF6; ISSN: 1381-5148
PB Elsevier Science B.V.
DT Journal
LA English
CC 35-5 (Chemistry of Synthetic High Polymers)
Section cross-reference(s): 76
AB The catalytic reaction of Co₂(**DAM**)₇Cl₄ with phenolic compds. such as phenol, p-cresol, o-cresol, m-cresol etc. in the presence of **hydrogen** peroxide gave C-O bonded oligomers (1-4). The oligomers were characterized by NMR, IR, elemental anal. The Mn values of 1-4 were 3417, 3494, 3438 and 3550, resp. The oligomers contain OH end groups.

The oligomer 1 has resistance profile such that it increases to 50.degree.C and then falls exponentially in the region of 50 to 170.degree.. The resistance of 2 and 4 are similar to semiconductor in the range of 30-180.degree.C. The oligomer 3 has a resistance profile such that it decreases from 30 to 100.degree.C and then increases from 100 to 140.degree. and once again falls in the region of 140-200.degree.. The oligomer 4 has an ESR signal at 3220 G and shows the presence of a free radical. The intensity of this signal increases ten times upon heating to 90.degree..

ST polyphenol oxidative polymn cobalt catalyst

IT Polymerization catalysts

(oxidative; polyphenols prep'd. by cobalt(II) catalyzed oxidative polymn.)

IT Electric resistance

(polyphenols prep'd. by cobalt(II) catalyzed oxidative polymn.)

IT Polyoxyphenylenes

RL: PRP (Properties); SPN (Synthetic preparation); PREP (Preparation)

(polyphenols prep'd. by cobalt(II) catalyzed oxidative polymn.)

IT 57-71-6D, cobalt trinuclear complex 7440-48-4D, Cobalt, diacetyl monoxime trinuclear complex

RL: CAT (Catalyst use); USES (Uses)

(polyphenols prep'd. by cobalt(II) catalyzed oxidative polymn.)

IT 25134-02-5P, o-Cresol homopolymer 25667-40-7P, Poly(1,4-phenylene oxide) 27073-41-2P, Phenol polymer 27289-33-4P 27289-34-5P, p-Cresol homopolymer

RL: PRP (Properties); SPN (Synthetic preparation); PREP (Preparation)

(polyphenols prep'd. by cobalt(II) catalyzed oxidative polymn.)

RE.CNT 38 THERE ARE 38 CITED REFERENCES AVAILABLE FOR THIS RECORD

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L26 ANSWER 8 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 2000:333039 HCPLUS

TI Low-temperature methane decomposition on various carbon-supported cobalt catalysts.

AU Zhang, Zhanguo; Haraguchi, K.; Yoshida, Tadashi

CS Energy Resources Division, Hokkaido National Industrial Research Institute, Sapporo, 062-8517, Japan

SO Book of Abstracts, 219th ACS National Meeting, San Francisco, CA, March 26-30, 2000 (2000), PETR-051 Publisher: American Chemical Society, Washington, D. C.

DT CODEN: 69CLAC

DT Conference; Meeting Abstract

LA English

AB It has been demonstrated in previous studies that both decompn. and hydrogenation reactions in the two-step non-oxidative homologation of methane were catalytic surface reaction and that the highest conversion of methane to C₂₊ hydrocarbons was reached at very low carbon coverages. Therefore, the high dispersion of active metal would be of great significance for increasing methane conversion and reducing the amt. of catalyst used. In order to develop such high dispersed catalysts, Co catalysts supported on various carbons with surface area from 80-2990m²/g were prepd. in this study. The characterization of catalysts was conducted by using BET adsorption, XRD anal., H₂-TPR and H₂-TPD techniques and the activity of the catalysts was tested by both CH₄-TPR and methane decompn. at a const. temp. of 450.degree.C, resp. The measurements of BET adsorption showed that carbons with surface area higher than 1000m²/g were of much high capacity for the loading of Co, in comparison with conventional SiO₂ support. Further, from the results of XRD anal., H₂-TPR and H₂-TPD, it was found that Co supported on these high surface area carbons was very well dispersed. On the other hand, the results of activity tests showed that the catalysts supported by these carbons with high surface areas, in comparison with Co/SiO₂, lowered the beginning temp. of hydrogen evolution in the CH₄-TPR and increased a great deal the initial prodn. rate of hydrogen in the decompn. reaction at 450.degree.C, resp. It is clear that the high activities of carbon supported Co catalysts resulted from their high dispersion. In addn., it was also found that the molar ratio of reacted CH₄ to produced H₂ was great than 0.5 within the initial 80 s in the methane decompn. at 450.degree.C. It suggests that CH_x(x<4) species, which are very important to the prodn. of C₂₊ in the second step hydrogenation reaction, were formed on the surface of Co metal.

L26 ANSWER 9 OF 52 HCPLUS COPYRIGHT 2002 ACS
AN 2000:168927 HCPLUS
DN 132:334835
TI Cobalt(II) diacetylmonoxime complex-catalysed oxidative coupling reactions of few aromatic compounds
AU Puzari, A.; Baruah, J. B.
CS Department of Chemistry, Indian Institute of Technology, Guwahati, India
SO J. Mol. Catal. A: Chem. (2000), 153(1-2), 1-6
CODEN: JMCCF2; ISSN: 1381-1169
PB Elsevier Science B.V.
DT Journal
LA English
CC 35-3 (Chemistry of Synthetic High Polymers)
Section cross-reference(s): 45
AB Catalytic amt. of a diacetylmonoxime cobalt(II) chloride complex [A] along with hydrogen peroxide reacts with toluene to give benzylalc., benzaldehyde and trace amt. of p-cresol and o-cresol. Similar catalytic reaction of aniline gives C-N-bonded oligomer having Mn and Mw of 1256 and 1892, resp. Mass spectra detd. by MALDI suggests the polyaniline to have highest mass from 14 units of aniline with CoCl₂ unit attached at the end. The oligomer thus formed have semiconducting properties. Similar reaction of phenol with cobalt(II) diacetylmonoxime (Co-DAM) complex in the presence of hydrogen peroxide gives C-O-bonded polyphenol. The reaction of 2-aminophenol with [A] and hydrogen peroxide leads to aggregate that has property of irreversible thermoelec. sensor. ESR suggests this property to originate from free radical present in the aggregate. Substituted aniline such as 2,4-dimethoxyaniline can also form low mol. aggregate from similar reaction that has semiconducting property in the range of 35.degree.C-180.degree.C.
ST cobalt diacetylmonoxime complex oxidative coupling catalyst; dimethoxyaniline cobalt diacetylmonoxime complex oxidative coupling; aminophenol cobalt diacetylmonoxime complex oxidative coupling
IT Oxidation catalysts
(cobalt diacetylmonoxime complex-catalyzed oxidative coupling reactions of arom. compds.)
IT Polyanilines
RL: SPN (Synthetic preparation); PREP (Preparation)
(cobalt diacetylmonoxime complex-catalyzed oxidative coupling reactions of arom. compds.)
IT Polymerization catalysts
(oxidative; cobalt diacetylmonoxime complex-catalyzed oxidative coupling reactions of arom. compds.)
IT Compd. [Co₂(MeCOCNOHMe)7Cl₄], IR 2916 cm⁻¹
RL: CAT (Catalyst use); SPN (Synthetic preparation);
PREP (Preparation); USES (Uses)
IT 57-71-6DP, Diacetylmonoxime, metal complexes 7440-48-4DP, Cobalt, diacetylmonoxime complex
RL: CAT (Catalyst use); SPN (Synthetic preparation);
PREP (Preparation); USES (Uses)
(cobalt diacetylmonoxime complex-catalyzed oxidative coupling reactions of arom. compds.)
IT 57-71-6, Diacetylmonoxime 108-88-3, reactions 7791-13-1,
Cobalt(II) chloride hexahydrate
RL: RCT (Reactant)
(cobalt diacetylmonoxime complex-catalyzed oxidative coupling reactions of arom. compds.)
IT 25233-30-1P, Polyaniline 25668-01-3P 27073-41-2P, Phenol polymer
267428-75-1P

RL: SPN (Synthetic preparation); PREP (Preparation)
(cobalt diacetylmonoxime complex-catalyzed oxidative coupling
reactions of arom. compds.)

RE.CNT 23 THERE ARE 23 CITED REFERENCES AVAILABLE FOR THIS RECORD
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L26 ANSWER 10 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD

AN 1999-321606 [27] WPIX

DNN N1999-241620 DNC C1999-094680

TI Oxidation catalyst for fuel battery - has a complex oxide of base metal, platinum group metal and rare earth metal which is dispersed in activated alumina and coated on integrated ceramic carrier.

DC E36 J04 L03 X16

PA (NSMO) NISSAN MOTOR CO LTD

CYC 1

PI JP 11114423 A 19990427 (199927)* 11p B01J023-63

ADT JP 11114423 A JP 1997-284066 19971016

PRAI JP 1997-284066 19971016

IC ICM B01J023-63

ICS C01B003-56; H01M008-06

AB JP 11114423 A UPAB: 19990714

NOVELTY - A complex oxide of a base metal such as copper containing platinum group metal and rare earth metal is dispersed in activated alumina and coated on the surface of a ceramic integrated carrier, to form the catalyst which selectively oxidizes carbon-monoxide (CO) in hydrogen gas.

DETAILED DESCRIPTION - INDEPENDENT CLAIMS are also included for the following: (i) Manufacture of oxidation catalyst which involves hydrothermal reaction of a carbonate with the intermediate formation of bicarbonate compound. (ii) method of removing CO from mixed gas containing hydrogen as the principal component and also oxygen and CO. CO is selectively converted to carbon dioxide (CO₂), which is separated from the gaseous mixture.

USE - For fuel battery used as electric power generator, car battery system etc.

ADVANTAGE - The reduction of the complex metal oxide in a catalyst to its metals is prevented. The amount of oxygen necessary for CO oxidation is supplied and the concentration of CO is monitored thereby selective oxidation of CO is performed. Time-dependent deterioration of the catalyst is reduced.

Dwg. 0/4

FS CPI EPI

FA AB; DCN

MC CPI: E11-Q01; E11-Q02; E31-A02; E31-D02; E31-N05B; J04-E04; L03-E04;
L03-H05

EPI: X16-C

L26 ANSWER 11 OF 52 HCAPLUS COPYRIGHT 2002 ACS

AN 1999:266087 HCAPLUS

DN 131:33660

TI Ru/SiO₂ catalysts prepared by the sol-gel method from Ru₃(CO)₁₂

AU Moggi, Pietro; Predieri, Giovanni; Di Silvestri, Fabio; Ferretti, Andrea
CS Dipartimento di Chimica Organica e Industriale, Universita di Parma,
Parma, 43100, Italy

SO Appl. Catal., A (1999), 182(2), 257-265

CODEN: ACAGE4; ISSN: 0926-860X

PB Elsevier Science B.V.

DT Journal

LA English

CC 51-11 (Fossil Fuels, Derivatives, and Related Products)

Section cross-reference(s): 45, 67

AB Ru/SiO₂ catalysts were prep'd. by sol-gel methods from Ru₃(CO)₁₂, or [Ru₃H(CO)₁₁]⁻, Si(OMe)₄ (TMOS) and H₂O, followed by thermal activation in helium up to 573 K and redn. with hydrogen. They were compared with Ru/SiO₂ corresponding material prep'd. via sol-gel from RuCl₃. The catalysts were characterized by FTIR spectra, surface area, metal dispersion and activity tests in the Fischer-Tropsch reaction at atm. pressure and CO/H₂ 1:1, in the 473-573 K temp. range. The use of ruthenium clusters as precursors afforded higher metal dispersions than RuCl₃, as well as higher catalytic activities in the hydrogenation of CO (particularly in the absence of alkali or halogen ions retained in the gel structure). TEM investigations on the Ru/SiO₂ material, prep'd. from Ru₃(CO)₁₂ by sol-gel, without promoters, showed the presence of very fine metal particles homogeneously embedded in the silica matrix, with dimensions ranging from 1-4 nm.

ST ruthenium catalyst precursor cluster sol gel prepn;
Fischer Tropsch catalyst ruthenium carbonyl sol gel method; carbon monoxide hydrogenation ruthenium silica catalyst

IT Fischer-Tropsch catalysts

Fischer-Tropsch reaction

Sol-gel processing

(Ru/SiO₂ catalysts prep'd. by the sol-gel method from Ru₃(CO)₁₂)

IT Hydrocarbons, preparation

RL: SPN (Synthetic preparation); PREP (Preparation)
(synthesis; Ru/SiO₂ catalysts prep'd. by the sol-gel method from Ru₃(CO)₁₂)

IT 7440-18-8, Ruthenium, uses 7631-86-9, Silica, uses

12125-01-8, Ammonium fluoride 15243-33-1, Ruthenium carbonyl
(ru3(co)12)

RL: CAT (Catalyst use); USES (Uses)
(Ru/SiO₂ catalysts prep'd. by the sol-gel method
from Ru₃(CO)₁₂)

IT 630-08-0, Carbon monoxide, reactions

RL: RCT (Reactant)
(hydrogenation; Ru/SiO₂ catalysts prep'd.
by the sol-gel method from Ru₃(CO)₁₂)

IT 681-84-5, Tetramethoxysilane 60496-59-5

RL: CAT (Catalyst use); USES (Uses)
(precursor; Ru/SiO₂ catalysts prep'd. by the sol-gel
method from Ru₃(CO)₁₂)

RE.CNT 40 THERE ARE 40 CITED REFERENCES AVAILABLE FOR THIS RECORD

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L26 ANSWER 12 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 1999:376154 HCPLUS

DN 131:107348

TI Hydrogenation catalysts prepared from amorphous alloys

with zirconium
AU Takahashi, Takeshige; Kai, Takami
CS Faculty of Engineering, Department of Applied Chemistry and Chemical
Engineering, Kagoshima University, Kagoshima, Japan
SO Mater. Sci. Eng., A (1999), A267(2), 207-213
CODEN: MSAPE3; ISSN: 0921-5093
PB Elsevier Science S.A.
DT Journal
LA English
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction
Mechanisms)
Section cross-reference(s): 56
AB The **hydrogenation** of benzene and carbon dioxide was carried out
over Zr-based **catalysts** prep'd. from amorphous alloys contg.
palladium, platinum, rhodium or nickel as the second metal.
Although the **hydrogenation** activity of the alloys was very low
in the virgin state, the activity gradually increased after oxidn. and
redn. in a **hydrogen** atm. Characterization of the alloys
revealed that the increase in surface area caused by oxidn. of zirconium
was responsible for the increase in **hydrogenation** activity.
When the amorphous alloys were oxidized in an oxygen atm., both the active
metals for the **hydrogenation** and zirconium in the alloys were
simultaneously oxidized. When the oxidized alloys were reduced in a
hydrogen atm., the oxides of the active metals were reduced to the
resp. metals, whereas zirconium remained in the oxide state. As a result,
palladium, platinum, rhodium or nickel were highly dispersed on
the produced zirconium oxide crystals. The principal advantage of a
catalyst prep'd. from an amorphous alloy is the very high
dispersion of **active metal** on the surface,
despite the high metal content. It is considered that amorphous alloys
contg. zirconium as the second metal are attractive **catalyst**
precursors for **hydrogenation**.
ST platinum zirconium amorphous alloy **hydrogenation**
catalyst; rhodium zirconium amorphous alloy **hydrogenation**
catalyst; nickel zirconium amorphous alloy **hydrogenation**
catalyst; benzene **hydrogenation** zirconium amorphous
alloy **catalyst**; carbon dioxide **hydrogenation** zirconium
amorphous alloy **catalyst**
IT Dispersion (of materials)
 Hydrogenation catalysts
 Oxidation
 Reduction
 Surface area
 (**hydrogenation catalysts** prep'd. from amorphous
 alloys with zirconium)
IT Metallic glasses
 RL: CAT (Catalyst use); PRP (Properties); USES (Uses)
 (**hydrogenation catalysts** prep'd. from amorphous
 alloys with zirconium)
IT 11146-73-9 62943-91-3 64568-51-0 75865-87-1 81160-75-0
115844-65-0 116294-11-2 230646-69-2 230646-70-5
 RL: CAT (Catalyst use); PRP (Properties); USES (Uses)
 (**hydrogenation catalysts** prep'd. from amorphous
 alloys with zirconium)
IT 71-43-2, Benzene, reactions 124-38-9, Carbon dioxide, reactions
1333-74-0, **Hydrogen**, reactions
 RL: RCT (Reactant)
 (**hydrogenation catalysts** prep'd. from amorphous
 alloys with zirconium)
RE.CNT 24 THERE ARE 24 CITED REFERENCES AVAILABLE FOR THIS RECORD

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L26 ANSWER 13 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 1999:53301 HCPLUS

DN 130:72055

TI Investigations of **palladium catalysts** on different carbon supports

AU Albers, P.; Burmeister, R.; Seibold, K.; Prescher, G.; Parker, S. F.; Ross, D. K.

CS Degussa AG, Hanau, D-63403, Germany

SO J. Catal. (1999), 181(1), 145-154

CODEN: JCTLA5; ISSN: 0021-9517

PB Academic Press

DT Journal

LA English

CC 67-1 (Catalysis, Reaction Kinetics, and **Inorganic** Reaction Mechanisms)

Section cross-reference(s): 66

AB **Pd-catalysts** (5 wt%) were prep'd. on various supports: powd. and pelletized activated carbon derived from pine wood or beech wood, carbon black, and graphitized carbon black. The supports were **palladized** in the original condition or after HCl-treatment. The **catalysts** were characterized with CO-chemisorption, transmission electron micrographs, and X-ray photoelectron spectrometry. The precious **metal dispersion** and the catalytic **activity** are improved by the acid treatment of activated carbon as well as carbon black supports. A selection of samples were also characterized by means of inelastic neutron scattering (INS) to study the **hydrogen-related** features of the acid modification of the carbon supports. Indications for different sensitivities of the final **Pd/C-catalysts** to catalytically driven hydrogasification effects of the supports and the impact of these effects on the precious metal dispersion and the formation of **Pd-hydrides** were obtained. The formation of mol.-like polyarom. structures was obsd. esp. on HCl-treated

activated carbon due to enhanced spillover of active **hydrogen** from the **Pd**-particles to adjacent sites on the support. The generation and accessability of these sites on the support seems to be promoted by the removal of ash and adsorbed species. (c) 1999 Academic Press.

ST palladium catalyst different carbon support
IT Hydrogenation catalysts
(for cinnamic acid; palladium catalysts on different carbon supports)
IT Catalyst supports
Catalysts
Dispersion (of materials)
Particle size distribution
(palladium catalysts on different carbon supports)
IT 7440-05-3, Palladium, uses 7440-44-0, Carbon, uses
RL: CAT (Catalyst use); PRP (Properties); USES (Uses)
(palladium catalysts on different carbon supports)
IT 7647-01-0, Hydrogen chloride, processes
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(palladium catalysts on different carbon supports)

RE.CNT 46 THERE ARE 46 CITED REFERENCES AVAILABLE FOR THIS RECORD
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L26 ANSWER 14 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 1998-362568 [31] WPIX
CR 1994-007502 [01]
DNC C1998-111553
TI Aqueous phase hydrogenation of organic materials - using a catalyst containing separate nickel and ruthenium phases dispersed on a support, giving high activity, long catalyst life and reduced agglomeration and sintering.
DC E19 H04 J04
IN ELLIOTT, D C; SEALOCK, L J; SEALOCK, J; SEALOCK, J L
PA (BATT) BATTELLE MEMORIAL INST
CYC 79
PI WO 9826869 A1 19980625 (199831)* EN 23p B01J023-89
RW: AT BE CH DE DK EA ES FI FR GB GH GM GR IE IT KE LS LU MC MW NL OA
PT SD SE SZ UG ZW
W: AL AM AT AU AZ BA BB BG BR BY CA CH CN CU CZ DE DK EE ES FI GB GE
GH HU IL IS JP KE KG KP KR KZ LC LK LR LS LT LU LV MD MG MK MN MW
MX NO NZ PL PT RO RU SD SE SG SI SK SL TJ TM TR TT UA UG UZ VN YU
ZW
AU 9857273 A 19980715 (199846) B01J023-89
US 5814112 A 19980929 (199846) C10J003-00
BR 9713764 A 20000201 (200023) B01J023-89
CN 1246077 A 20000301 (200029) B01J023-89
MX 9905782 A1 19991001 (200103) B01J023-89
AU 727906 B 20010104 (200107) B01J023-89
ADT WO 9826869 A1 WO 1997-US24176 19971218; AU 9857273 A AU 1998-57273
19971218; US 5814112 A Cont of US 1992-893701 19920605, CIP of US
1994-227892 19940415, US 1996-775337 19961219; BR 9713764 A BR 1997-13764
19971218, WO 1997-US24176 19971218; CN 1246077 A CN 1997-181764 19971218;
MX 9905782 A1 MX 1999-5782 19990618; AU 727906 B AU 1998-57273 19971218
FDT AU 9857273 A Based on WO 9826869; US 5814112 A CIP of US 5616154; BR
9713764 A Based on WO 9826869; AU 727906 B Previous Publ. AU 9857273,
Based on WO 9826869
PRAI US 1996-775337 19961219; US 1992-893701 19920605; US 1994-227892
19940415
IC ICM B01J023-89; C10J003-00
ICS C01B003-32; C02F001-02
AB WO 9826869 A UPAB: 20010202
Aqueous phase hydrogenation of organic materials comprises: (a) providing a liquid reactant mixture containing liquid water and the organic material in a pressure reactor; (b) adding a particulate catalyst, each particle comprising a porous support containing (i) a deposited reduced nickel metal catalyst phase first phase; and (ii) an added ruthenium metal as a separate and distinct dispersed second phase, in which the amount of ruthenium phase is sufficient to resist agglomeration or sintering of the nickel phase; and (c) maintaining the liquid reactant mixture at 150-350 deg. C and a hydrogen overpressure.
USE - The hydrogenation reaction comprises saturation of multiple bonds, reduction of oxygen functional groups (e.g. ketones,

aldehydes, lactones and carboxylic acids), hydrogenolysis (e.g. reduction of alcohols and phenolics), scission of ether linkages, corresponding reactions of organically bound nitrogen, hydrocracking and/or methane synthesis (claimed), e.g. steam reforming to produce a product gas containing methane, carbon dioxide and **hydrogen**; sorbitol conversion to ethylene glycol, propylene glycol, and glycerol; levulinic acid conversion to gamma -valerolactone; 1,4-pentanediol and methyl tetrahydrofuran, xylitol conversion to ethylene glycol and glycerol; succinic acid conversion to gamma -butyrolactone, 1,4-butanediol and tetrahydrofuran. Potential feedstocks include waste water streams, e.g. 0.5% hexamethylenediamine in water from nylon manufacture, or 2-3% mixed phenols in water from resin manufacture that would be treated primarily for waste destruction; and cheese whey, peat, or high-moisture biomass feedstocks for energy recovery systems.

ADVANTAGE - The porous support is stable during use, remaining physically intact and chemically inert without coming apart or disintegrating (claimed). The **ruthenium metal dispersed** phase increases **activity** compared to the use of nickel alone and increases the effective lifetime of the **catalyst**. The process can treat liquid organic materials having wide ranges of organic concentration, giving conversions of, e.g. 99.9%.

Dwg.0/2

FS CPI
FA AB; DCN
MC CPI: E10-E04; E10-J02B3; E10-J02D; H04-B03; H04-E04; H04-E08; H04-F02;
H04-F02E; J04-E01; N02-C; N02-E01

L26 ANSWER 15 OF 52 HCPLUS COPYRIGHT 2002 ACS
AN 1998:293872 HCPLUS
DN 128:309766
TI A comparative evaluation of high-temperature membrane systems for catalytic processing
AU Gobina, Edward; Hou, Kaihu; Hughes, Ronald
CS Dep. Chem. Eng., Univ. Salford, Salford, M5 4WT, UK
SO Chem. Eng. Commun. (1998), 166, 157-181
CODEN: CEGCAK; ISSN: 0098-6445
PB Gordon & Breach Science Publishers
DT Journal
LA English
CC 47-9 (Apparatus and Plant Equipment)
Section cross-reference(s): 67
AB The relative performance of various composite membranes incorporated in high-temp. membrane reactors was evaluated by math. simulation. Two membrane categories (porous and dense) and 6 types of composite membrane systems (**Pd/Ag**, polyimide, silica, inert porous and **Ru**-dispersed porous silicalite zeolites, and C mol. sieve) were compared. The special case of the **Pd/Ag** composite system having imperfections (pinholes and cracks) is also considered. The industrially important reaction of ethylbenzene dehydrogenation over promoted **iron oxide catalyst** using rate const. values from the literature was studied as the model reaction. While the dense systems possess higher performance levels at lower membrane thicknesses, the porous composite systems esp. those contg. highly **dispersed active metal** particles within their micropores have advantages due to their significantly higher contact surface to vol. ratio. The study has also confirmed that high H permselectivity is a key factor in detg. reactor performance in terms of conversion enhancement.
ST simulation composite membrane catalytic processing; porous dense membrane catalytic processing evaluation; ethylbenzene dehydrogenation composite membrane comparative evaluation; **hydrogen** permselectivity

ethylbenzene dehydrogenation composite membrane

IT Molecular sieves
(carbon, membrane; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT Catalysis
Dehydrogenation **catalysts**
Mass transfer
Physicochemical simulation
(comparative evaluation of high-temp. membrane systems for catalytic processing)

IT Membranes (nonbiological)
(composite; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT Polyimides, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(membrane; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT Permselectivity
(of **hydrogen** in comparative evaluation of high-temp. membrane systems for catalytic processing)

IT Silicalites (zeolites)
RL: TEM (Technical or engineered material use); USES (Uses)
(porous membrane; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT 7440-18-8, **Ruthenium**, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(-dispersed porous membrane; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT 7440-22-4, **Silver**, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(-**palladium** membrane; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT 7440-05-3, **Palladium**, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(-silver membrane; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT 100-41-4, **Ethylbenzene**, reactions
RL: RCT (Reactant)
(comparative evaluation of high-temp. membrane systems for catalytic dehydrogenation of)

IT 7631-86-9, **Silica**, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(membrane; comparative evaluation of high-temp. membrane systems for catalytic processing)

IT 1333-74-0, **Hydrogen**, processes
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(permselectivity in comparative evaluation of high-temp. membrane systems for catalytic processing)

L26 ANSWER 16 OF 52 HCAPLUS COPYRIGHT 2002 ACS

AN 1997:283335 HCAPLUS

DN 126:295391

TI **Metal dispersion and catalytic activity in Pd catalysts** for methane combustion

AU Dalla-Betta, R. A.; Loffler, D. G.; Magno, S.

CS Catalytica Combustion Systems, Inc., Mountain View, CA, 94043, USA

SO Prepr. - Am. Chem. Soc., Div. Pet. Chem. (1997), 42(1), 163-165

CODEN: ACPCAT; ISSN: 0569-3799

PB American Chemical Society, Division of Petroleum Chemistry

DT Journal

LA English
CC 51-12 (Fossil Fuels, Derivatives, and Related Products)
Section cross-reference(s): 67
AB The light-off temp. of zirconia-supported palladium catalysts for methane combustion can be calcd. from the hydrogen adsorption uptake of the wash coat powder. This correlation constitutes a useful tool for designing combustion catalysts for gas turbine applications.
ST zirconia palladium catalyst methane combustion; metal dispersion palladium combustion catalyst; lightoff temp zirconia palladium combustion catalyst
IT Chemisorption
Combustion catalysts
(metal dispersion and catalytic activity in Pd catalysts for methane combustion)
IT 1314-23-4, Zirconium oxide (ZrO₂), uses 7440-05-3, Palladium, uses
RL: CAT (Catalyst use); USES (Uses)
(metal dispersion and catalytic activity in Pd catalysts for methane combustion)
IT 74-82-8, Methane, uses
RL: NUU (Other use, unclassified); USES (Uses)
(metal dispersion and catalytic activity in Pd catalysts for methane combustion)

L26 ANSWER 17 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1997:162979 HCAPLUS
TI Metal dispersion and catalytic activity in Pd catalysts for methane combustion
AU Dalla Betta, R. A.; Loffler, D. G.; Magno, S.
CS Catalytica Combustion Systems, Inc., Mountain View, CA, 94043, USA
SO Book of Abstracts, 213th ACS National Meeting, San Francisco, April 13-17 (1997), PETR-058 Publisher: American Chemical Society, Washington, D. C.
CODEN: 64AOAA
DT Conference; Meeting Abstract
LA English
AB Characterization of Pd-based catalysts for methane combustion frequently include measuring metal dispersion using either H₂ or CO chemisorption. In spite of uncertainties derived from the formation of palladium oxides under reaction conditions, those techniques generally result in reproducible results. Yet, the correlation of catalytic activity with phys. properties of the catalyst is not well established. Literature reports suggest that, under reaction conditions, bulk palladium oxide is formed. Oxidn. breaks apart the metal crystallites, so that most of the oxide participates in the reaction. This process may be reversed when the sample is reduced prior to measuring hydrogen uptake. We have prep'd. a series of Pd-based catalysts for methane combustion with different metal loading. The catalytic activity of the samples was correlated with metal dispersion measured using CO and H₂ chemisorption. The results suggest a linear correlation between hydrogen chemisorption and catalyst activity for samples with relatively high metal dispersion. This correlation seems to break down for poorly dispersed samples.

L26 ANSWER 18 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1997:33798 HCAPLUS
DN 126:62577
TI Method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal

IN Huffman, Gerald P.; Taghieei, M. Mehdi
 PA University of Kentucky Research Foundation, USA
 SO Can. Pat. Appl., 20 pp.
 CODEN: CPXXEB
 DT Patent
 LA English
 IC ICM C10G001-08
 CC 51-24 (Fossil Fuels, Derivatives, and Related Products)
 Section cross-reference(s): 60

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	CA 2171803	AA	19960916	CA 1996-2171803	19960314
PRAI	US 1995-403985		19950315		

AB A method for the liquefaction of waste plastic in a reaction vessel, comprising the steps of: charging the reaction vessel with waste plastic and a slurring agent wherein the wt. ratio of slurring agent to waste plastic is substantially between 0 and 5:1; charging the reaction vessel with a processing gas, selected from a group consisting of hydrogen, nitrogen and mixts. thereof, to a pressure between 15-3,000 psig cold; contacting the waste plastic, slurring agent and processing gas in the reaction vessel with a catalyst selected from a group consisting of zeolites, metal impregnated zeolites, ion exchanged zeolites, hydrogenated zeolites, silica-alumina, metal-doped silica-alumina, activated clays, activated fly ashes; highly dispersed metal oxides, oxyhydroxides and ferrihydrites; sulfated metal oxides, oxyhydroxides and ferrihydrites; secondary metal substituted oxides, oxyhydroxides and ferrihydrites; acid treated metal oxides, oxyhydroxides and ferrihydrites; and any mixts. thereof; heating the waste plastic, slurring agent, processing gas and catalyst in the reaction vessel to a temp. of substantially 350-500.degree. for a residence time of substantially 15-120 min; and recovering resulting liquefied products.

ST liquefaction waste plastic coal

IT Fly ash

RL: CAT (Catalyst use); USES (Uses)
 (activated; method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)

IT Waste plastics

(method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)

IT Activated clays

Aluminosilicates, uses

Zeolite HZSM-5

Zeolites (synthetic), uses

RL: CAT (Catalyst use); USES (Uses)
 (method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)

IT Coal, processes

RL: PEP (Physical, engineering or chemical process); PROC (Process)
 (method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)

IT Naphtha

Rubber, uses

RL: TEM (Technical or engineered material use); USES (Uses)
 (slurrying agent; method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)

IT Petroleum refining residues

(vacuum distn., slurrying agent; method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)

IT Petroleum refining
(waste plastic liquefaction; method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)
IT 1333-74-0, Hydrogen, uses 7727-37-9, Nitrogen, uses
RL: TEM (Technical or engineered material use); USES (Uses)
(method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)
IT 119-64-2, 1,2,3,4-Tetrahydronaphthalene 613-31-0, Dihydroanthracene
20330-24-9, Hexahydriopyrene
RL: TEM (Technical or engineered material use); USES (Uses)
(slurrying agent; method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)
IT 624-92-0, Dimethyl disulfide 7704-34-9, Sulfur, reactions 7783-06-4,
Dihydrogen sulfide, reactions
RL: RCT (Reactant)
(sulfur donor; method of liquefaction of waste plastics and of co-liquefaction of waste plastics and coal)

L26 ANSWER 19 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1996:11995 HCAPLUS
DN 124:127959
TI Preparation of **hydrogenation catalysts** from amorphous alloys
AU Takahashi, Takeshige; Kai, Takami
CS Dep. Applied Chem. and Chem. Eng., Kagoshima Univ., Kogoshima, 890, Japan
SO Kagaku Kogaku Ronbunshu (1995), 21(6), 961-71
CODEN: KKRBAW; ISSN: 0386-216X
DT Journal
LA Japanese
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)
Section cross-reference(s): 25, 56
AB Hydrogenation of benzene and carbon dioxide was carried out over catalysts prep'd. from amorphous alloys contg. zirconium. The hydrogenation activity of the alloys was very low in the virgin state, but the activity gradually increased in the order of oxidn., redn. in a hydrogen atm. and **hydrogenation**. Characterization of the amorphous alloys revealed that the increase in surface area caused by oxidn. of zirconium was responsible for the increase in catalytic activity. When the alloys were oxidized in an oxygen atm. the active metal atoms for **hydrogenation** and zirconium in the alloys were simultaneously oxidized. When the oxidized alloys were treated in a **hydrogen** atm., **palladium**, platinum and nickel oxides were reduced to the resp. metal atoms, whereas zirconium oxide remained. As a result, **palladium**, platinum or nickel **catalysts** supported on zirconium oxide were prep'd. from the amorphous alloys. The principal advantage of a **catalyst** prep'd. from an amorphous ally was high dispersion of active metal on the surface, despite the high content of metal. Amorphous alloys contg. zirconium are considered to be an attractive **catalyst** precursor for **hydrogenation**.
ST **hydrogenation catalyst** prep'n zirconium amorphous alloy
IT Hydrogenation catalysts
(prep'n. of **hydrogenation catalysts** from amorphous alloys)
IT Metallic glasses
RL: CAT (Catalyst use); USES (Uses)
(prep'n. of **hydrogenation catalysts** from amorphous alloys)
IT 71-43-2, Benzene, reactions 124-38-9, Carbon dioxide, reactions

RL: RCT (Reactant)
(hydrogenation of; prepn. of **hydrogenation catalysts** from amorphous alloys)

IT 1314-23-4, Zirconia, uses 7440-02-0, Nickel, uses 7440-05-3,
Palladium, uses 7440-06-4, Platinum, uses 11146-74-0
56293-95-9 72211-92-8 75865-90-6 80740-85-8 82283-10-1
107529-41-9 173104-36-4 173104-37-5 173104-38-6 173104-39-7
173104-40-0

RL: CAT (Catalyst use); USES (Uses)
(prepn. of **hydrogenation catalysts** from amorphous alloys)

L26 ANSWER 20 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1995:799707 HCAPLUS
DN 123:209779

TI Preparation of highly dispersed supported catalysts by ultrasound

AU Bianchi, C. L.; Carli, R.; Fontaneto, C.; Ragaini, V.
CS Department Physical Chemistry and Electrochemistry, University Milan,
Milan, 20133, Italy
SO Stud. Surf. Sci. Catal. (1995), 91(Preparation of Catalysts VI), 195-1100
CODEN: SSCTDM; ISSN: 0167-2991

DT Journal
LA English
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)

AB Palladium supported on alumina or active carbon catalysts were prepd. using ultrasound during the prepn. steps. A large increase in the **metal dispersion** and in the catalytic activity of the samples, tested during the redn. of acetophenone with flowing **hydrogen**, was found.

ST highly dispersed supported catalyst ultrasound;
palladium dispersed supported catalyst prepn ultrasound

IT Reduction catalysts
(for acetophenone; prepn. of highly dispersed supported catalysts by ultrasound)

IT Catalysts and Catalysis
Sound and Ultrasound
(prepn. of highly dispersed supported catalysts by ultrasound)

IT 1344-28-1, Alumina, uses 7440-05-3, Palladium, uses 7440-44-0, Carbon, uses

RL: CAT (Catalyst use); PEP (Physical, engineering or chemical process); PROC (Process); USES (Uses)
(prepn. of highly dispersed supported catalysts by ultrasound)

L26 ANSWER 21 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1994:461453 HCAPLUS
DN 121:61453

TI hydriding properties of Mg-based **hydrogen** storage materials prepared chemically from a homogeneous phase

AU Imamura, Hayao; Nakamura, Masahiro
CS Faculty of Engineering, Yamaguchi University, Ube, 755, Japan
SO Z. Phys. Chem. (Munich) (1994), 183(1-2), 157-62
CODEN: ZPCFAX; ISSN: 0044-3336

DT Journal
LA English
CC 52-3 (Electrochemical, Radiational, and Thermal Energy Technology)
Section cross-reference(s): 56, 78

AB Dispersion of Mg atoms into NH₃ at 77 K using a metal vapor technique was used to produce a homogeneous soln. of dissolved Mg metal in liq. ammonia. The Mg/NH₃ soln. was used to prep. novel Mg-based H storage materials. Materials prep'd. include: active, small Mg metals crystd. out of the soln. in the presence or absence of catalytically active Ni powders (Mg-Ni or Mg). The Mg metal was highly **dispersed** on an **active** carbon (AC) support with high surface area by impregnating the AC with the soln. of Mg/NH₃. By using AC on which active Ni, Ru or Pt metals had been highly dispersed, the prepn. of samples was further extended to include binary systems (Mg-Ni/AC, Mg-Ru /AC or Mg-Pt/AC). Samples characteristics were correlated with the prepn. methods. These Mg-based compds. were extremely active toward H absorption.

ST hydrogen absorption magnesium ammonia dispersion

IT Absorption

(of **hydrogen**, by magnesium dispersed in ammonia, role of metals in)

IT 1333-74-0, Hydrogen, properties

RL: PRP (Properties)

(absorption of, by magnesium/ammonia dispersion, hydriding characteristics in, for storage)

IT 7440-44-0, Carbon, uses

RL: USES (Uses)

(activated, **hydrogen** absorption by magnesium dispersion in ammonia in presence of)

IT 7440-02-0, Nickel, uses 7440-06-4, Platinum, uses 7440-18-8, Ruthenium, uses

RL: CAT (Catalyst use); USES (Uses)

(catalyst, **hydrogen** absorption by magnesium in presence of, for storage)

IT 7664-41-7, Ammonia, uses

RL: USES (Uses)

(dispersant, magnesium hydriding characteristics in, for storage)

IT 7439-95-4, Magnesium, properties

RL: PRP (Properties)

(**hydrogen** absorption by ammonia dispersion of, hydriding characteristics in, for storage)

L26 ANSWER 22 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 1994:16395 HCPLUS

DN 120:16395

TI Influence of preparation method on the characteristics of nickel/sepiolite **catalysts**

AU Anderson, James A.; Daza, Loreto; Fierro, Jose Luis G.; Rodrigo, M. Theresa

CS Inst. Catal. Petroleoquim., CSIC, Madrid, 28049, Spain

SO J. Chem. Soc., Faraday Trans. (1993), 89(19), 3651-7

CODEN: JCFTEV; ISSN: 0956-5000

DT Journal

LA English

CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)

AB Two series of nickel **catalysts** supported on sepiolite were prep'd. by impregnation and pptn. procedures, and the resultant materials characterized in their reduced states. FTIR of adsorbed CO and XPS were used to study the surfaces of the materials and the degree of redn. was calcd. gravimetrically. The metal dispersion/particle size was calcd. from **hydrogen** adsorption isotherms at 298 K, and the **hydrogenation** of benzene used as a test reaction of the metal sites. No correlation between **activity** and **metal**

dispersion was found and turnover frequencies indicate an apparent structure sensitivity of the reaction in addn. to an apparent influence of the prepn. method. However, similarity between the form of plots of turnover frequency-particle size and degree of redn.-particle size, suggests that the apparent structure sensitivity is the result of the presence of unreduced Ni in the catalyst surfaces. The presence of this unreduced Ni acts as a diluent in the Ni surface, thereby reducing the size of Ni ensembles.

- ST nickel sepiolite catalyst prepn method characteristics;
hydrogenation catalyst nickel sepiolite
IT Catalysts and Catalysis
 Hydrogenation catalysts
 (nickel-sepiolite, influence of prepn. method on characteristics of)
IT Particle size
Reduction
 (of nickel-sepiolite catalysts, influence of prepn. method on)
IT 63800-37-3, Sepiolite
RL: CAT (Catalyst use); USES (Uses)
 (catalyst from nickel and, influence of prepn. method on characteristics of)
IT 7440-02-0, Nickel, uses
RL: CAT (Catalyst use); USES (Uses)
 (catalyst from sepiolite and, influence of prepn. method on characteristics of)

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- L26 ANSWER 23 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1995:956253 HCAPLUS
DN 123:345294
TI Chemical role of Athabasca oil sand bitumen for high conversion of coal-heavy oil coprocessing and evaluation of highly dispersed metal catalysts
AU Higashine, Jun; Kido, Akihiro; Takahashi, Katsuyuki; Miyake, Mikio; Nomura, Masakatsu
CS Faculty Engineering, Osaka University, Suita, 565, Japan
SO Proc. - Annu. Int. Pittsburgh Coal Conf. (1993), 10th, 241-5
CODEN: PICNE4; ISSN: 1075-7961
DT Journal
LA English
CC 51-14 (Fossil Fuels, Derivatives, and Related Products)
AB A mixt. of coal and pure solvent such as 1-methylnaphthalene (MN), biphenyl (BP), Decaline (DC), and n-eicosane (EC) with 1 to 3 ratio was coprocessed in a 70 mL autoclave at 420 .degree.C for 1 h in the presence of H₂ (50 kg/cm² at room temp.) and sulfided Ni-Mo/Al₂O₃ catalyst. The resulting reactivity was compared with that of coal with Athabasca oil sand bitumen (AOB). As for the run with Wandoan subbituminous coal, HS (hexane sol. fraction) yield was high with AOB compared with the run in the presence of DC and EC, while the reverse was obsd. with HS yield from the run with Taiheiyo subbituminous coal. These results were explained by referring to the two features of solvent, that is a dispersant of radicals produced and a hydrogen shuttler.
The activity of highly dispersed metal catalyst from Fe(CO)₅, Co₂(CO)₈, and Ru₃(CO)₁₂ for this coprocessing reaction (Wandoan coal and AOB) was also examd.
ST oil sand bitumen coal coprocessing catalyst
IT Coal liquefaction
 (chem. role of Athabasca oil sand bitumen for high conversion of coal-heavy oil coprocessing)
IT Coal liquefaction catalysts

Tar oils

(chem. role of Athabasca oil sand bitumen for high conversion of coal-heavy oil coprocessing and evaluation of highly dispersed metal catalysts)

IT 10210-68-1 13463-40-6, Iron pentacarbonyl 15243-33-1,
Ruthenium carbonyl (Ru₃(CO)₁₂)
RL: CAT (Catalyst use); USES (Uses)
(chem. role of Athabasca oil sand bitumen for high conversion of coal-heavy oil coprocessing and evaluation of highly dispersed metal catalysts)

L26 ANSWER 24 OF 52 HCAPLUS COPYRIGHT 2002 ACS

AN 1994:201488 HCAPLUS

DN 120:201488

TI Carbon monoxide/hydrogen titration studies on Ru
:Mn/alumina supported bimetallic system for the determination of metal surface area

AU Hussain, S. Tajammul

CS A. Q. Khan Res. Lab., Rawalpindi, Pak.

SO J. Chem. Soc. Pak. (1993), 15(4), 234-7

CODEN: JCSPDF; ISSN: 0253-5106

DT Journal

LA English

CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)

Section cross-reference(s): 66

AB The CO/H₂ titrn. method was used to det. the metal surface area on Ru:Mn/Al₂O₃ supported system. This method was based on the assumption that CO is selectively adsorbed to form a monolayer on the metal surface but is not adsorbed on the support surface and that all the adsorbed CO is hydrogenated to form CH₄ selectively. The data indicate that Mn covers the surface of Ru and hence the no. of active sites responsible for the prodn. of CH₄ is blocked by the addn. of Mn. A continuous decrease in the dispersion of the active metal is discussed in terms of ensemble affect generated by the presence of Mn on the surface of Ru.

ST ruthenium manganese alumina catalyst metal surface;
carbon monoxide titrn ruthenium manganese catalyst;
hydrogen titrn ruthenium manganese catalyst

IT Adsorption

(carbon monoxide-hydrogen titrn. for metal surface area detn.
for ruthenium-manganese/alumina catalysts)

IT Catalysts and Catalysis

(ruthenium-manganese/alumina, carbon monoxide-hydrogen titrn. for metal surface area detn. for)

IT 7440-18-8, Ruthenium, uses

RL: CAT (Catalyst use); USES (Uses)
(catalysts from manganese and, supported on alumina, carbon monoxide-hydrogen titrn. for metal surface area detn. for)

IT 7439-96-5, Manganese, uses

RL: CAT (Catalyst use); USES (Uses)
(catalysts from ruthenium and, supported on alumina, carbon monoxide-hydrogen titrn. for metal surface area detn. for)

IT 1344-28-1, Alumina, uses

RL: CAT (Catalyst use); USES (Uses)
(catalysts from ruthenium-manganese on, carbon monoxide-hydrogen titrn. for metal surface area detn. for)

IT 1333-74-0, Hydrogen, reactions

RL: RCT (Reactant)
(titrn. of carbon monoxide and, for metal surface area detn. for
ruthenium-manganese/alumina catalysts)
IT 630-08-0, Carbon monoxide, reactions
RL: RCT (Reactant)
(titrn. of **hydrogen** and, for metal surface area detn. for
ruthenium-manganese/alumina catalysts)

L26 ANSWER 25 OF 52 HCPLUS COPYRIGHT 2002 ACS
AN 1992:538664 HCPLUS
DN 117:138664
TI Process for controlling degree of dispersion of active
component of metal/carrier solid catalyst
IN Matsuura, Ikuya; Yoshida, Yasushi; Takayasu, Osamu; Nitta, Kuniaki
PA Ube Industries, Ltd., Japan
SO U.S., 6 pp.
CODEN: USXXAM
DT Patent
LA English
IC ICM B01J023-58
 ICS B01J023-78
NCL 502328000
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction
Mechanisms)
FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	US 5137863	A	19920811	US 1991-775086	19911011
	JP 04150950	A2	19920525	JP 1990-275371	19901016
	JP 04331704	A2	19921119	JP 1991-188392	19910426
	JP 05208801	A2	19930820	JP 1991-188391	19910426
PRAI	JP 1990-275371		19901016		
	JP 1991-188391		19910426		
	JP 1991-188392		19910426		

AB Disclosed is a process for controlling the degree of dispersion of a catalytically active component in the prepn. of a metal/carrier solid catalyst, which comprises dry-blending an ultrafine single-crystal carrier material and an active component metal material, molding the blend, heating and maintaining the molded body in an inert gas at a temp. higher than and close to the m.p. of the active component metal material, and heat-treating the molded body in the same atm. gas at a temp. higher than the decompr. temp. of the active component metal material.
ST metal dispersion control supported catalyst
IT Catalysts and Catalysis
(control of metal dispersion on supported)
IT 1309-48-4, Magnesium oxide, uses
RL: CAT (Catalyst use); USES (Uses)
(catalyst from precious metal and, control of metal
dispersion on)
IT 7440-02-0, Nickel, miscellaneous 7440-05-3, Palladium,
miscellaneous 7440-06-4, Platinum, miscellaneous 7440-16-6, Rhodium,
miscellaneous 7440-18-8, Ruthenium, miscellaneous
RL: MSC (Miscellaneous)
(control of dispersion of, in supported catalyst)
IT 74-82-8, Methane, reactions
RL: RCT (Reactant)
(conversion of, to carbon monoxide and hydrogen, precious
metal-magnesium oxide catalyst for)
IT 630-08-0P, Carbon monoxide, preparation 1333-74-0P, Hydrogen,
preparation

RL: SPN (Synthetic preparation); PREP (Preparation)
 (prepn. of, from methane, precious metal-magnesium oxide catalyst for)

L26 ANSWER 26 OF 52 HCPLUS COPYRIGHT 2002 ACS
 AN 1992:242645 HCPLUS
 DN 116:242645
 TI Process for controlling degree of dispersion of active component of metal/carrier solid catalyst
 IN Matsuura, Ikuya; Takayasu, Osamu; Yoshida, Yasushi; Nitta, Kuniaki
 PA Ube Industries, Ltd., Japan
 SO Eur. Pat. Appl., 11 pp.
 CODEN: EPXXDW
 DT Patent
 LA English
 IC ICM B01J037-00
 ICS B01J037-08
 CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)

FAN.CNT 2

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	EP 481689	A1	19920422	EP 1991-309361	19911010
	R: DE, DK, FR, GB, IT, NL				
	JP 04150950	A2	19920525	JP 1990-275371	19901016
PRAI	JP 1990-275371		19901016		

AB Disclosed is a process for controlling the degree of dispersion of a catalyst active component in the prepn. of a metal/carrier solid catalyst, which comprises dry-blending an ultrafine single-crystal carrier material and an active component metal material, molding the blend, heating and maintaining the molded body in an inert gas at a temp. higher than and close to the m.p. of the active component metal material, and heat-treating the molded body in the same atm. gas at a temp. higher than the decomprn. temp. of the active component metla material.

ST metal dispersion control supported catalyst

IT Catalysts and Catalysis

(supported, control of metal dispersion of)

IT 1309-48-4, Magnesium oxide, properties

RL: PRP (Properties)

(control of dispersion of metals on, for catalysts)

IT 7440-02-0, Nickel, properties 7440-05-3, Palladium, properties

7440-06-4, Platinum, properties 7440-16-6, Rhodium, properties

7440-18-8, Ruthenium, properties

RL: PRP (Properties)

(control of dispersion of, for supported catalysts)

IT 74-82-8, Methane, reactions

RL: RCT (Reactant)

(conversion of, to carbon monoxide and hydrogen, control of metal dispersion on magnesium oxide of catalysts for)

IT 630-08-0P, Carbon monoxide, preparation 1333-74-0P, Hydrogen, preparation

RL: SPN (Synthetic preparation); PREP (Preparation)

(prepn. of, from methane, control of metal dispersion on magnesium oxide of catalysts for)

L26 ANSWER 27 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
 AN 1991-233811 [32] WPIX
 DNC C1991-101688

TI Alcohol contg. methanol a a main component, prepn. - comprising reacting hydrogen with carbon mon oxide in presence of hydrophobic

catalyst.

DC E17 J04
PA (MITK) MITSUI TOATSU CHEM INC
CYC 1
PI JP 03151337 A 19910627 (199132)* 3p
JP 2764080 B2 19980611 (199828) 3p C07C031-04
ADT JP 03151337 A JP 1989-287593 19891106; JP 2764080 B2 JP 1989-287593
19891106
FDT JP 2764080 B2 Previous Publ. JP 03151337
PRAI JP 1989-287593 19891106
IC B01J027-12; B01J033-00; C07B061-00; C07C029-15; C07C031-04
ICM C07C031-04
ICS B01J027-12; B01J031-24; B01J033-00; C07C029-15; C07C029-153
ICA C07B061-00
AB JP 03151337 A UPAB: 19930928
Alcohol, mainly methanol (MC), is prep'd. by reaction of H₂ (hydrogen) with CO₂ or with a mixed gas of CO₂ and CO in solvent immiscible with water or aq. MC in the presence of a hydrophobic catalyst used for methanol synthesis. The heterogeneous catalyst comprises an active component on hydrophobic carrier (e.g. polytetrafluoroethylene, black lead fluoride, etc.) or a coated known methanol synthesis catalyst with hydrophobic cpd. (e.g. hydrophobically (PbF₂) treated Cu-Zn) are used. Cu, Pd, or Pt, etc. is used as a metal on the carrier. Zn or Cr, etc. are also carried for stabilisation and high dispersion of the active metal.

USE/ADVANTAGE - MC is effectively produced from CO₂ and water due to no adsorption of water on the surface of the catalyst.

0/0

FS CPI
FA AB; DCN
MC CPI: E10-E04E1; J04-E01; N02-D01; N02-F02; N03-D; N03-F

L26 ANSWER 28 OF 52 HCPLUS COPYRIGHT 2002 ACS
AN 1991:172241 HCPLUS
DN 114:172241
TI Study of highly dispersed rhodium cluster derived catalysts
AU Chen, Yu.; Chen, Yaoqiang; Jiang, Xinrong; Xie, Hue
CS Sichuan Univ., Chengdu, Peop. Rep. China
SO Fenzi Cuihua (1990), 4(2), 148-55
CODEN: FECUEN
DT Journal
LA Chinese
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)
Section cross-reference(s): 51
AB Highly dispersed Rh₄/Al₂O₃ and Rh₂/Al₂O₃ catalysts were prep'd. and characterized by H₂ chemisorption and IR spectra of adsorbed CO. The activity and selectivity of the catalysts under a series of H₂/CO ratios were investigated in a micro-flow reactor. The results showed that there was a high degree of metal dispersion and a uniform metal particle size on both catalysts. In the process of heating in vacuum or in a CO atm., only the twin chemisorbed CO type appeared, neither the spectra of the linear type nor that of a bridge type was obsd. The spectra of twin chemisorbed type disappeared in H₂. The activity of the 2 catalysts increased with temp. and with higher H₂/CO ratios. Both catalysts had high methanation activity and selectivity, but under the same conditions. Rh₄/Al₂O₃ had a higher activity than Rh₂/Al₂O₃.

ST rhodium cluster alumina **catalyst** prepn characterization;
 dispersion rhodium alumina **catalyst**; **hydrogen**
 chemisorption rhodium alumina **catalyst**; carbon monoxide
hydrogenation rhodium alumina **catalyst**

IT Adsorbed substances
 (carbon monoxide, on rhodium-alumina **catalysts**, IR spectra
 of, effect of **hydrogen** presence on)

IT Chemisorption
 (of **hydrogen**, on rhodium-alumina **catalysts**,
 dispersion in relation to)

IT Particle size
 (of rhodium support on alumina, catalytic activity in relation to)

IT Catalysts and Catalysis
 (rhodium clusters supported on alumina, prepn. and characterization of)

IT Hydrogenation catalysts
 (rhodium-alumina, for carbon monoxide, effect of **metal**
 dispersion on **activity** and selectivity of)

IT 630-08-0, Carbon monoxide, properties
 RL: PRP (Properties)
 (adsorbed, on rhodium cluster support on alumina, IR spectra of, effect
 of **hydrogen** addn. on)

IT 7440-16-6, Rhodium, uses and miscellaneous
 RL: CAT (Catalyst use); USES (Uses)
 (catalysts from alumina and, **metal**
 dispersion of, **activity** in **hydrogenation** of
 carbon monoxide in relation to)

IT 1333-74-0, Hydrogen, reactions
 RL: RCT (Reactant)
 (chemisorption of, on rhodium cluster support on alumina, dispersion in
 relation to)

L26 ANSWER 29 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
 AN 1989-250073 [35] WPIX
 DNC C1989-111368
 TI Converting polysaccharide(s) to poly hydric alcohol(s) - by
hydrogenation in presence of **catalyst** comprising
 supported metal e.g. **ruthenium** and acidic solid.
 DC E17
 IN HINNEKENS, H; JACOBS, P
 PA (LABO) FINA RES SA; (SYNFINA-N) SYNFINA-OLEOFINA SA
 CYC 16
 PI EP 329923 A 19890830 (198935)* EN 10p
 R: AT BE CH DE ES FR GB GR IT LI LU NL SE
 DK 8900784 A 19890823 (198943)
 JP 01268653 A 19891026 (198949)
 US 4950812 A 19900821 (199036)
 EP 329923 B1 19930421 (199316) EN 11p C07C031-26
 R: AT BE CH DE ES FR GB GR IT LI LU NL SE
 DE 3880477 G 19930527 (199322) C07C031-26
 ES 2054877 T3 19940816 (199434) C07C031-26
 ADT EP 329923 A EP 1988-870023 19880222; JP 01268653 A JP 1989-39495 19890221;
 US 4950812 A US 1989-313946 19890222; EP 329923 B1 EP 1988-870023
 19880222; DE 3880477 G DE 1988-3880477 19880222, EP 1988-870023 19880222;
 ES 2054877 T3 EP 1988-870023 19880222
 FDT DE 3880477 G Based on EP 329923; ES 2054877 T3 Based on EP 329923
 PRAI EP 1988-870023 19880222
 REP 1.Jnl.Ref; US 2609399
 IC ICM C07C031-26
 ICS B01J029-06; C07C029-14; C13K013-00
 AB EP 329923 A UPAB: 19960417

Polysaccharides are converted to polyhydric alcohols by hydrogenation at high pressure and temp. in the presence of a catalyst consisting of : i) a supported metal, (being either ruthenium, copper, nickel or cobalt) dispersed on the support as to adsorb more than 0.58 molecules CO per metal atom and ii) a solid with sufficient acid functions so that the rate constant of hydrolysis of sucrose, k_1 , on the catalyst is greater than 70% of the rate constant of hydrogenation of glucose, k_2 , on the catalyst.

The supported metal is prepared by ion exchange or the decomposition of an ammine complex salt. The support may be identical to the solid, such as a synthetic acidic zeolite. The reaction temp. ranges from 348 to 523 K.

ADVANTAGE - This is a single step process combining hydrolysis and hydrogenation and results in high purity hexitols.

0/0

Dwg.0/0

FS CPI

FA AB; DCN

MC CPI: E10-A07; N02; N06-E

L26 ANSWER 30 OF 52 HCAPLUS COPYRIGHT 2002 ACS

AN 1990:105785 HCAPLUS

DN 112:105785

TI Thermal decomposition of metal carbonyls on oxide supports containing surface hydrides: a route to highly dispersed metal catalysts with unusual properties

AU Lisitsyn, A. S.; Golovin, A. V.; Chuvilin, A. L.; Kuznetsov, V. L.; Romanenko, A. V.; Danilyuk, A. F.; Ermakov, Yu. I.

CS Inst. Catal., Novosibirsk, 630090, USSR

SO Appl. Catal. (1989), 55(2), 235-58

CODEN: APCADI; ISSN: 0166-9834

DT Journal

LA English

CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)

Section cross-reference(s): 51, 66, 77

AB A method of prepn. of fairly dispersed metal catalysts on oxide supports is described. The Co, Fe, Fe-Co, and Rh carbonyls, when chemisorbed on SiO₂ with anchored hydrides of Ti or Al and then subjected to thermal decompr., produce very small species consisting of several up to tens of metal atoms. Thermal decompr. of a heteronuclear Fe-Co cluster leads to formation of bimetallic Fe-Co particles. The Co, Fe, and Co-Fe samples so obtained possess unusual magnetic properties. On varying temp. and strength of the magnetic field, their magnetization alters in a manner specific for super-paramagnetic particles, but there is no signal of ferromagnetic resonance in ESR spectra. With increased loading of the supported Fe and Co carbonyls, both chemisorbed and weakly bound carbonyl complexes are on the surface of the support. Larger metallic crystallites are much smaller (.apprx.1 nm) than those obtained on non-modified SiO₂. The metal particles in the samples show high stability to sintering under vacuum but grow in size under conditions of CO hydrogenation (1.5-2 nm). Carbide formation presumably takes place on the Fe catalyst during CO hydrogenation. The catalytic properties in the CO + H₂ reaction are affected by both the size of metal particles in the catalysts and the chem. nature of the support. ST metal carbonyl thermal decompr silica support; iron silicon

catalyst prepn carbonyl; cobalt silica catalyst
prepн carbonyl; ruthenium silica catalyst prepн
carbonyl; particle size catalyst prepн carbonyl
IT Hydrogenation catalysts
(cobalt and/or iron or rhodium supported on silica,
prepд. by thermal decompn. of metal carbonyls,
dispersion and activity of)
IT Particle size
(of cobalt and/or iron or rhodium supported on
silica, prepд. by thermal decompn. of metal carbonyls, catalytic
activity in relation to)
IT Magnetic induction and Magnetization
(of cobalt and/or iron supported on silica, prepд.
by thermal decompn. of metal carbonyls, temp. dependence of)
IT 7439-89-6P, Iron, uses and miscellaneous 7440-16-6P, Rhodium,
uses and miscellaneous 7440-48-4P, Cobalt, uses and
miscellaneous
RL: CAT (Catalyst use); PREP (Preparation); USES
(Uses)
(catalysts from silica and, prepд. by thermal decompn. of
adsorbed carbonyl, dispersion of)
IT 7429-90-5, Aluminum, uses and miscellaneous 7440-32-6, Titanium, uses
and miscellaneous
RL: USES (Uses)
(silica catalyst supports modified by, dispersion of metal
formed by thermal decompn. of adsorbed metal carbonyl in relation to)

L26 ANSWER 31 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1989:102507 HCAPLUS
DN 110:102507
TI Effect of the support on the catalytic activity of a platinum-
rhenium catalyst
AU Cid, R.; Pecchi, G.; Reyes, P.
CS Fac. Cienc., Univ. Concepcion, Concepcion, Chile
SO Bol. Soc. Chil. Quim. (1988), 33(4), 229-37
CODEN: BOCQAX; ISSN: 0366-1644
DT Journal
LA Spanish
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction
Mechanisms)
Section cross-reference(s): 22
AB A series of Pt-Re catalysts supported on SiO₂, Al₂O₃,
and NaY zeolite with different Pt/Re ratios were prepд. Several
methods were used to characterize the solids: chem. anal., a
potentiometric method to measure the surface acidity, O₂-H₂
titrns. to evaluate the metal dispersion. As a test reaction, the
decompn. of methylcyclohexane in H was studied. The nature of the support
had a great influence on the surface acidity, metal
dispersion, activity, and stability of the
catalysts. The PtRe/Al₂O₃ catalysts are the most acidic
and are better dispersed; however, they show a very fast deactivation.
ST platinum rhenium supported catalyst acidity
dispersion; silica platinum rhenium catalyst acidity
dispersion; alumina platinum rhenium catalyst acidity
dispersion; zeolite platinum rhenium catalyst acidity
dispersion; methylcyclohexane decomprn supported platinum rhenium
catalyst
IT Chemisorbed substances
(hydrogen, on supported platinum-rhenium
catalysts, oxygen titrn. of, metal dispersion in relation to)

IT Catalysts and Catalysis
(platinum-rhenium-alumina or -silica or -zeolite, acidity and metal dispersion in activities of)

IT Decomposition catalysts
(supported platinum-rhenium, for Me cyclohexane, effect of support material on activity of)

IT Zeolites, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(NaY, catalysts from platinum and rhenium and, acidity and dispersion and activity of, effect of compn. on)

IT 7440-15-5, Rhenium, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalysts from platinum and various supports and, effect of support material on acidity and dispersion and activity of)

IT 7440-06-4, Platinum, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalysts from rhenium and various supports and, effect of support material on acidity and dispersion and activity of)

IT 1333-74-0, Hydrogen, uses and miscellaneous
RL: USES (Uses)
(chemisorbed, on supported platinum-rhenium catalysts, titrn. of, by oxygen, metal dispersion in relation to)

IT 7782-44-7, Oxygen, uses and miscellaneous
RL: USES (Uses)
(titrn. by, of hydrogen chemisorbed on supported platinum-rhenium catalysts, metal dispersion in relation to)

IT 1335-30-4
RL: RCT (Reactant)
(zeolites, NaY, catalysts from platinum and rhenium and, acidity and dispersion and activity of, effect of compn. on)

L26 ANSWER 32 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1988:516788 HCAPLUS
DN 109:116788
TI Effect of hydrogen chloride on the activity and dispersion of platinum and palladium supported on aluminum oxide
AU Anan'in, V. N.; Pryakhina, N. Yu.; Gerasimova, L. V.; Sen'kov, G. M.; Trokhimets, A. I.
CS Inst. Fiz.-Org. Khim., Minsk, USSR
SO Vestsi Akad. Navuk BSSR, Ser. Khim. Navuk (1988), (3), 30-3
CODEN: VBSKAK; ISSN: 0002-3590
DT Journal
LA Russian
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)
AB A method is described for prep. catalysts with enhanced activity and Pt or Pd dispersion which involves impregnation after Al₂O₃ surface treatment by gaseous HCl. Catalyst activities were detd. for cyclohexane dehydrogenation in a flow reactor.
ST platinum alumina catalyst prep impregnation; palladium alumina catalyst dehydrogenation; hydrogen chloride treatment alumina surface
IT Catalysts and Catalysis
Dehydrogenation catalysts
(palladium or platinum, on alumina supports, hydrogen chloride pretreatment of support in relation to enhanced activity of)

IT 7647-01-0, Hydrogen chloride, uses and miscellaneous
RL: USES (Uses)
(alumina catalyst support pretreatment with, enhanced activity and metal dispersion in relation

to)

IT 1344-28-1, Alumina, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalyst support, hydrogen chloride pretreatment
effects on metal dispersion on)

IT 7440-05-3P, Palladium, uses and miscellaneous 7440-06-4P,
Platinum, uses and miscellaneous
RL: CAT (Catalyst use); PREP (Preparation); USES
(Uses)
(catalysts, on alumina supports, prepd. by impregnation,
hydrogen chloride pretreatment of support effect on)

L26 ANSWER 33 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1988:535765 HCAPLUS
DN 109:135765
TI Effect of palladium concentration on the state and activity of
palladium/.gamma.-alumina catalysts
AU Parvulescu, Vasile; Sandulescu, Ioan; Grecu, Nicoleta; Contescu, Cristian;
Nicolescu, Ioan V.
CS Natl. Inst. Chem., Bucharest, Rom.
SO Rev. Roum. Chim. (1988), 33(1), 3-11
CODEN: RRCHAX; ISSN: 0035-3930
DT Journal
LA English
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction
Mechanisms)
Section cross-reference(s): 22
AB H chemisorption and EPR measurements on Pd/.gamma.-Al₂O₃
catalysts were performed to det. the effect of Pd concn.
on the state and the dispersion of the metal on the support. From the EPR
spectra, Pd+1 clusters were found. The EPR and dispersion
results are discussed in connection with the activity measurements.
ST palladium alumina catalyst dispersion;
hydrogen chemisorption palladium alumina
catalyst; cluster palladium alumina catalyst
IT Chemisorption
(of hydrogen, on palladium-alumina
catalysts, particle size in relation to)
IT Particle size
(of palladium supported on alumina, effect of
palladium loading on, activity in relation to)
IT Hydrogenation catalysts
(palladium-alumina, effect of metal loading on
dispersion and activity of)
IT 7440-05-3, Palladium, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalysts from alumina and, effect of palladium
concn. on dispersion and activity of)
IT 1333-74-0, Hydrogen, reactions
RL: RCT (Reactant)
(chemisorption of, on palladium-alumina catalysts,
particle size in relation to)

L26 ANSWER 34 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1988:97265 HCAPLUS
DN 108:97265
TI Catalyst for preparation of trichlorosilane from silicon
tetrachloride
IN Morimoto, Shiro
PA Osaka Titanium Co., Ltd., Japan

SO Jpn. Kokai Tokkyo Koho, 5 pp.

CODEN: JKXXAF

DT Patent

LA Japanese

IC ICM C01B033-107

ICS B01J027-128

CC 49-8 (Industrial Inorganic Chemicals)

FAN.CNT 1

	PATENT NO.	KIND	DATE	APPLICATION NO.	DATE
PI	JP 62235205	A2	19871015	JP 1986-79486	19860407
OS	CASREACT 108:97265				
AB	Si(s) and a highly active H-redn. chloride-dispersed transition metal catalyst which is supported on spherical, gas-permeable naphthalene-processed petroleum pitch are placed in a reaction chamber and gaseous SiCl ₄ -H ₂ is fed into the chamber to give SiHCl ₃ . Effective conversion of SiCl ₄ into SiHCl ₃ is achieved with short contact time. Aq. CoCl ₂ soln. was mixed with Kureha Beads BAC-MP (support) and dried to give a catalyst precursor carrying Co chloride which was treated 5-6 h at 500.degree. with Si(s) in H ₂ (g) flow. A mixt. of 0.186 mmol/min SiCl ₄ and 0.421 mmol/min H ₂ was fed into a reaction chamber contg. the catalyst (heated to 500.degree.) to give 88.7 mol% SiCl ₄ and 11.0 mol% SiHCl ₃ in 25.0 s residence time.				
ST	silicon chloride hydride prepn catalyst; chlorosilane prepn catalyst				
IT	7440-02-0, Nickel, uses and miscellaneous 7440-05-3, Palladium, uses and miscellaneous 7440-06-4, Platinum, uses and miscellaneous 7440-16-6, Rhodium, uses and miscellaneous 7440-48-4D, Cobalt, supported on pitch beads 7440-50-8, Copper, uses and miscellaneous RL: CAT (Catalyst use); USES (Uses) (catalyst, in redn. of silicon tetrachloride)				
IT	10025-78-2P, Trichlorosilane RL: PREP (Preparation) (prepn. of, from silicon tetrachloride, supported catalyst for)				
IT	10026-04-7, Silicon tetrachloride RL: RCT (Reactant) (redn. of, supported catalyst for, in prepn. of trichlorosilane)				
L26	ANSWER 35 OF 52 HCPLUS COPYRIGHT 2002 ACS				
AN	1987:215844 HCPLUS				
DN	106:215844				
TI	The effect of catalyst preparation on catalytic activity. II. The design of nickel/alumina catalysts prepared by wet impregnation				
AU	Huang, Y. J.; Schwarz, J. A.				
CS	Dep. Chem. Eng. Mater. Sci., Syracuse Univ., Syracuse, NY, 13244, USA				
SO	Appl. Catal. (1987), 30(2), 255-63				
	CODEN: APCADI; ISSN: 0166-9834				
DT	Journal				
LA	English				
CC	45-4 (Industrial Organic Chemicals, Leather, Fats, and Waxes) Section cross-reference(s): 67				
AB	Design equations are presented for Ni/Al ₂ O ₃ catalysts prep'd. by wet impregnation from nickel nitrate soln. in contact with a .gamma.-Al ₂ O ₃ support. The metal dispersion, activity for Cl-3 formation from 3:1 H ₂ -CO, and the carbon deposited during reaction are predictable based solely on the properties				

of the electrolytes from which these **catalysts** are formed.
ST nickel **catalyst hydrogenation** carbon monoxide;
hydrocarbon prodn carbon monoxide
IT Kinetics of **hydrogenation**
(of carbon monoxide in prepn. of hydrocarbons, **catalyst**
prep. method effect on)
IT **Hydrogenation catalysts**
(supported, nickel, for carbon monoxide in prepn. of hydrocarbons,
activity of, prep. method effect on)
IT Hydrocarbons, preparation
RL: IMF (Industrial manufacture); PREP (Preparation)
(C1-3, prepn. of, by **hydrogenation** of carbon monoxide,
activity of supported nickel **catalysts** for)
IT 7440-02-0, Nickel, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(**catalysts**, on alumina, for **hydrogenation** of carbon
monoxide to hydrocarbons, activity of, method of prep. effect on)
IT 7440-44-0, Carbon, uses and miscellaneous
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(deposition of, on alumina-supported nickel **catalysts** for
conversion of carbon monoxide to hydrocarbons, **catalyst**
prep. method effect on)
IT 630-08-0, Carbon monoxide, reactions
RL: RCT (Reactant)
(**hydrogenation** of, in prodn. of hydrocarbons, supported
nickel **catalysts** for, effect of **catalyst** prep. on
activity of)

L26 ANSWER 36 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1987:73608 HCAPLUS
DN 106:73608
TI Characterization of carbon-supported **iron catalysts**
prepared from **iron** pentacarbonyl
AU Guerrero-Ruiz, A.; Lopez-Gonzalez, J. de D.; Mata-Arjona, A.;
Romero-Sanches, V.; Rodriguez-Ramos, I.
CS Dep. Inorg. Chem., Univ. Granada, Granada, Spain
SO Adsorpt. Sci. Technol. (1987), 3(1), 33-40
CODEN: ASTEEZ; ISSN: 0263-6174
DT Journal
LA English
CC 67-1 (Catalysis, Reaction Kinetics, and **Inorganic** Reaction
Mechanisms)
AB **Fe catalysts** were prep'd. by adsorption of **Fe**
(CO)5 on activated C. After redn. in H2 flow at
different temps., the metal dispersions and av. crystallite sizes were
detd. The **catalysts** were characterized by H2 and
CO chemisorption at 373 K, x-ray diffraction line broadening, and
TEM. The **Fe** crystallites prep'd. by this method present a narrow
particle size distribution and have high metal dispersions. CO
chemisorption at 373 K is a good technique to study **Fe** supported
catalysts.
ST iron carbon **catalyst** dispersion particle size; redn
dispersion iron carbon **catalyst**; chemisorption carbon
monoxide **hydrogen** iron **catalyst**
IT **Catalysts** and Catalysis
(iron-carbon, metal dispersion and particle-size distribution
of, effect of redn. temp. on)
IT Chemisorption
(of **hydrogen** and carbon monoxide on iron-carbon
catalyst, surface structure in relation to)

IT Surface structure
 (of iron supported on carbon, effect of redn. temp. on)
 IT Reduction
 (of iron supported on carbon, metal dispersion in particle size distribution in relation to temp. of)
 IT Particle size
 (of iron, supported on carbon, effect of redn. temp. on)
 IT 7440-44-0, Carbon, uses and miscellaneous
 RL: USES (Uses)
 (activated, catalyst from iron and, dispersion and particle size distribution of, effect of redn. temp. on)
 IT 7439-89-6, Iron, uses and miscellaneous
 RL: CAT (Catalyst use); USES (Uses)
 (catalyst from activated carbon and, metal dispersion and particle size distribution of, effect of redn. temp. on)
 IT 630-08-0, Carbon monoxide, reactions 1333-74-0, Hydrogen, reactions
 RL: RCT (Reactant)
 (chemisorption of, on iron-carbon catalyst, surface structure in relation to)

L26 ANSWER 37 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
 AN 1986-035593 [05] WPIX
 DNC C1986-015109

TI Hydrogenation of cpds. contg. alkenyl, alkynyl, aryl, cyano, imino - carbonyl or carboxyl gps. using active carbon catalyst comprising dispersion of transition metal in porous carbon matrix.

DC E19 E35 J04

IN WENNERBERG, A N

PA (STAD) STANDARD OIL CO INDIANA

CYC 1

PI US 4564702 A 19860114 (198605)* 12p

ADT US 4564702 A US 1983-489407 19830428

PRAI US 1983-489407 19830428

IC C07C063-04

AB US 4564702 A UPAB: 19930922

Hydrogenation of an alkenyl, alkynyl, aryl, cyano, amino, carbonyl or carboxyl gp. in 1 or more organic cpds. comprises contact with H₂ at ambient temp. to 260 deg.C using an active C catalyst having a cage-like structure and a BET surface area of at least 800 sq.m. per g. and a bulk density of at least 0.1g per cc which comprises a substantially uniform dispersion of a transition metal, transition metal-contg. material, or both in a porous C matrix, wherein the total concn. of dispersed metal and dispersed metal-contg. material is 0.001-30wt.%, calculated as the elemental metal and based on the wt. of the catalyst.

The catalyst is pref. by: (a) forming a uniform co-crystallite of a precursor of the transition metal or transition-metal contg. material and of a C precursor; (b) forming a uniform powdered mixt. of the co-crystallite and inorganic solids comprising an alkali metal hydroxide; (c) pyrolysing the powdered mixt. in an inert atmos. at 400-980 deg.C to form the C(matrix) having the metal or metal material dispersed therein; and (d) sepg. unreacted inorganic material and inorganic reaction prods., other than the dispersed metal or metal-contg. material, from the C(matrix). Alternatively step (a) may be replaced by: (i) forming a C precursor contg. the metal or metal-contg. material by the chemical reaction in soln. of: (a) a soluble C precursor having at least 1 anionic gp. chemically bound thereto; and (b) a soluble cation or a

soluble cationic complex of a transition metal, and (ii) pptg. and drying the metal-contg. C precursor.

ADVANTAGE - Substantially improved selectivity and efficiency in the H-transfer reactions is achieved.

O/O

FS CPI

FA AB

MC CPI: E10-B04B; E10-E04; E10-J02A; E10-J02D; E31-N04; E34-C01; E35; J04-E01; N01-C02; N02; ~~N03~~; N04-A

L26 ANSWER 38 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 1986-212450 [33] WPIX

DNC C1986-091452

TI Activation and modification of supported palladium catalyst - esp. for hydrogenation by simple and flexible calcination in gas contg. oxygen.

DC J04

IN DEHNER, R; GLOMBIK, A; HEILMANN, A; KOTTER, J; LIESKE, H; QUECK, S; SCHUBERT, R; VOLTER, J; GRASSHOFF, E; HATTWIG, M; KECK, M; KOETTER, J; LANGE, J; LUNAU, J; MEYE, H; SCHOEDEL, R; STUERTZ, H; VOELTER, J

PA (VELW) VEB LEUNA-WERKE ULRICH W; (VELW) LEUNA-WERKE GMBH

CYC 1

PI DD 234798 A 19860416 (198633)* 12p
DD 234798 B5 19950727 (199540)

B01J023-44

ADT DD 234798 A DD 1985-273400 19850220; DD 234798 B5 DD 1985-273400 19850220

PRAI DD 1985-273400 19850220

IC B01J023-44; B01J037-48

ICM B01J023-44

ICS B01J037-08; B01J037-14; B01J037-48

AB DD 234798 A UPAB: 19930922

Activation and modification of supported Pd catalysts, opt. contg. other components, by calcination is carried out by heating for up to 48 h at 400-1373 K in a gas contg. 0.1-100 vol% O₂ at a pressure of at least 0.1 MPa, followed by redn. if necessary.

The amt. of gas w.r.t. the vol. of catalyst is up to 2500, pref. up to 1000 v/vh; the duration of treatment max. 24 hr., esp. 0.5-12 hr; and the treatment temp. over 1150 K at an O₂ partial pressure of at least 0.1 MPa and a stationary gas atmos.; 600-1200 K in a gas stream contg. under 15 vol% O₂; 750-1150 K in a gas atmos. contg. less than 5 vol% O₂ in conjunction with N₂, other inert gases, H₂, gas mixts. contg. H₂ and combustion gases contg. impurities and with a low O₂ content; 873-1173 K in an atmos. contg. at least 15, esp. 10-100 vol% O₂; and 1023-1150 K in air, air-O₂ mixts. or pure O₂.

USE/ADVANTAGE - The catalyst is useful for selective hydrogenation. The process has high capacity, is flexible and simple and allows the hydrogenation activity and metal dispersity to be increased or reduced as required.

O/O

FS CPI

FA AB

MC CPI: J04-E04B; N02-F02

L26 ANSWER 39 OF 52 HCPLUS COPYRIGHT 2002 ACS
AN 1986:224589 HCPLUS

DN 104:224589

TI Colloidal palladium supported on chelate resin containing iminodiacetic acid groups as hydrogenation catalyst

AU Hirai, Hidefumi; Komatsuzaki, Shigeru; Toshima, Naoki

CS Fac. Eng., Univ. Tokyo, Tokyo, 113, Japan

SO J. Macromol. Sci., Chem. (1986), A23(8), 933-54

CODEN: JMCHBD; ISSN: 0022-233X

DT Journal

LA English

CC 24-4 (Alicyclic Compounds)

Section cross-reference(s): 67

AB Colloidal Pd supported on a chelate resin contg. iminodiacetic acid groups was prep'd. by refluxing the Pd chelate resin in methanol-water. Using the resin-supported colloidal Pd as a catalyst, cyclopentadiene was hydrogenated to cyclopentene in 97.1% selectivity at 100% conversion of cyclopentadiene under 1 atm of hydrogen in methanol at 30.degree.C. Finely dispersed metal particles ranging from 10 to 60 .ANG. in diam. were obsd. in the resin by electron microscopy. Both x-ray microanal. for Pd and elution anal. of Pd ion with an aq. soln. of disodium salt of EDTA demonstrated the existence of large amts. of Pd ion complexes in the resin. The amt. of Pd metal in the resin was estd. to be .apprx.5% of the total Pd. Since the resin, after removal of most of the ionic Pd, exhibited almost the same catalytic activity as before, it was concluded that the finely dispersed metal particles are the active species in the catalyst.

ST colloid palladium chelate resin catalyst
hydrogenation; polymer iminodiacetate palladium
catalyst

IT Hydrogenation catalysts
(palladium colloidal supported on chelating resin contg.
iminodiacetic acid groups, for cyclopentadiene)

IT 7440-05-3, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalyst, colloidal, on chelate resin contg. iminodiacetic
acid groups, for hydrogenation)

IT 142-73-4
RL: PRP (Properties)
(chelate resin contg. groups of, palladium colloidal
catalyst support from)

IT 542-92-7, reactions
RL: RCT (Reactant)
(hydrogenation of, palladium supported on chelated
resin as catalysts for)

L26 ANSWER 40 OF 52 HCAPLUS COPYRIGHT 2002 ACS

AN 1986:231309 HCAPLUS

DN 104:231309

TI Platinum catalysts supported on activated carbons. I.
Preparation and characterization

AU Rodriguez-Reinoso, F.; Rodriguez-Ramos, I.; Moreno-Castilla, C.;
Guerrero-Ruiz, A.; Lopez-Gonzalez, J. D.

CS Dep. Quim. Inorg., Univ. Alicante, Alicante, Spain

SO J. Catal. (1986), 99(1), 171-83

CODEN: JCTLA5; ISSN: 0021-9517

DT Journal

LA English

CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)

AB Several Pt catalysts supported on activated carbons (manufd. from olive stones and almond shells) were prep'd. with both H₂PtCl₆.6H₂O and [Pt(NH₃)₄]Cl₂ as metal precursor and by using different methods. Once reduced, the supported catalysts were characterized by H and CO chemisorption as well as x-ray diffraction and TEM. The effect of redn. conditions on metal dispersion was studied and correlated with

the surface properties of the supports. Pore sizes of 9-11 nm are necessary for obtaining a high Pt dispersion. To increase the dispersion of the catalysts prep'd. from [Pt(NH₃)₄]Cl₂, a treatment with He prior to redn. of the catalysts in H is essential to avoid the formation of an unstable hydride which leads to agglomeration of the Pt particles.

- ST platinum catalyst activated carbon prep'n; pore size platinum dispersion carbon
- IT Chemisorption
(of hydrogen and carbon monoxide, on platinum-carbon catalyst, surface structure in relation to)
- IT Particle size
(of platinum, on activated carbon, effect of prep'n. method on)
- IT Reduction catalysts
(of platinum-carbon, metal dispersion in relation to)
- IT 7440-44-0P, uses and miscellaneous
RL: PREP (Preparation)
(activated, catalysts from platinum and, metal dispersion in relation to prep'n. of)
- IT 7440-06-4, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalysts from activated carbon and, metal dispersion in relation to prep'n. method of)
- IT 630-08-0, reactions 1333-74-0, reactions
RL: RCT (Reactant)
(chemisorption of, on platinum-carbon catalysts, particle size and porosity in relation to)

- L26 ANSWER 41 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 1985-190420 [32] WPIX
DNN N1985-142823 DNC C1985-083161
TI Testing granular supported metal catalyst for activity and dispersity - by combined reaction in U-shaped reactor and pulsed carbon mon oxide chemisorption.
DC J04 S03
IN KECK, M; LAMBRECHT, W; MEYE, H; NEUBAUER, H D; NEUMANN, U; SCHODEL, R; SCHUBERT, R
PA (VELW) VEB LEUNA-WERKE ULRICH W
CYC 1
PI DD 220826 A 19850410 (198532)* 5p
DD 220826 B 19880525 (198841)
ADT DD 220826 A DD 1984-259769 19840201
PRAI DD 1984-259769 19840201
IC B01J037-00; G01N007-00; G01N031-10
AB DD 220826 A UPAB: 19930925
Combined determin. of the activity and dispersity of granular supported metal catalysts (I) in the 273-873 K temp. range at normal pressure up to pressures of 50 MPa is carried out with a U-shaped flow reactor and selective gas chemisorption by the pulsed flow method. 100-2000 mg (I) are heated to the redn. in the activation gas at a given pressure in a reactor in a radiant furnace. After activation, (I) is cooled to the reaction temp. and charged with the reaction gas. After at least 30 min, the reactivity is tested, then (I) is purged with at least 20 000 v/v H₂ for 10-180, pref. 30 min in the same reactor and hence brought to a temp. at least 10 K higher than the reaction temp. but max. to the activation temp. After this, it is shock cooled to 273 K and the flow is adjusted to 4-6 l/h. Then 0.2-0.4 cc CO pulses are passed over (I) at normal pressure at 4-6 min intervals until the metal surface is satd. Throughout the process, the gas flows first through the (I) charge and only then through the frit of the reactor.

USE/ADVANTAGE - The method is reliable, saves time, energy and material and is useful for quality control.

1/1

FS CPI EPI

FA AB

MC CPI: J04-C03; J04-E01; N06-D

EPI: S03-E12

L26 ANSWER 42 OF 52 HCPLUS COPYRIGHT 2002 ACS

AN 1986:156548 HCPLUS

DN 104:156548

TI Comparative study of nickel **catalysts** supported on X- and Y-zeolites

AU Chien, Shu Hua; Lu, Kuang Lieh; Huang, Hsiao Wen

CS Inst. Chem., Acad. Sin., Taipei, 115, Taiwan

SO J. Chin. Chem. Soc. (Taipei) (1985), 32(3), 309-15

CODEN: JCCTAC; ISSN: 0009-4536

DT Journal

LA English

CC 67-2 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)

Section cross-reference(s): 51

AB A comparative study of CO/H₂ reactions over Ni **catalysts** supported on X- and Y-zeolites, prep'd. by incipient-wetness impregnation (W) and ion-exchange (E) methods was done. The **catalysts** were characterized by surface area and volumetric chemisorption measurements, temp.-programmed redn. profiles, and x-ray diffraction anal. The turnover frequency decreases in the order: NiX(W) > NiY(W) > NiY(E) > NiX(E). The **metal dispersion** and **activity** of the **catalyst** are affected mainly by the support materials, the prep'n. methods, and the activation procedures. The high activity of NiX(W) is due to its higher reducibility, smaller Ni cryst. size, and better Ni dispersion. The low activity of NiX(E) is attributed to the poor ion-exchange of Ni with NaX, the poor reducibility, the large cryst. size, and the poor Ni dispersion.

ST nickel zeolite **catalyst hydrogenation** carbon monoxide; surface structure nickel zeolite **catalyst**; reducibility nickel zeolite **catalyst** activity; dispersion nickel zeolite **catalyst** activity

IT **Hydrogenation catalysts**

(nickel-zeolites, for carbon monoxide, activity of, effect of prep'n. method on)

IT Chemisorption

(of **hydrogen**, on nickel-zeolite **catalysts**, activity in relation to)

IT Surface structure

(of nickel-zeolite **catalysts**, activity in relation to)

IT Zeolites, uses and miscellaneous

RL: **CAT (Catalyst use)**; USES (Uses)
(X, **catalysts** from nickel and, for **hydrogenation** of carbon monoxide, effect of prep'n. method on activity of)

IT Zeolites, uses and miscellaneous

RL: **CAT (Catalyst use)**; USES (Uses)
(Y, **catalysts** from nickel and, for **hydrogenation** of carbon monoxide, effect of prep'n. method on activity of)

IT 7440-02-0, uses and miscellaneous

RL: **CAT (Catalyst use)**; USES (Uses)
(**catalysts** from zeolites and, for **hydrogenation** of carbon monoxide, effect of support material and prep'n. method on activity of)

IT 1333-74-0, reactions
RL: RCT (Reactant)
(chemisorption of, on nickel-zeolite **catalysts**, activity in relation to)

IT 630-08-0, reactions
RL: RCT (Reactant)
(hydrogenation of, on nickel-zeolite **catalysts**)

L26 ANSWER 43 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1986:24797 HCAPLUS
DN 104:24797
TI Solvated metal atom dispersed (SMAD) **catalysts**
. Highly active bimetallic systems
AU Imizu, Yuzo; Klabunde, Kenneth J.
CS Dep. Chem., Kansas State Univ., Manhattan, KS, 66506, USA
SO Chem. Ind. (Dekker) (1985), 22(Catal. Org. React.), 225-50
CODEN: CHEIDI; ISSN: 0737-8025
DT Journal
LA English
CC 67-1 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)
Section cross-reference(s): 22
AB New bimetallic catalytic materials were prepd. by the solvated metal atom dispersion (SMAD) procedure. A series of 0 valent bimetallic composites Co-Cr/SiO₂, Co-Mn/SiO₂, and Co-Fe/SiO₂ were prepd. and studied in catalytic test reactions and by H₂ chemisorption. An increase in activity for the Co/SiO₂ catalyst was found when Mn was inserted into the Co particles; the increase being sufficient to create a catalyst active enough to **hydrogenate** 1-butene at -60.degree. in a diffusion-controlled process. Selectivity/activity studies coupled with H₂ chemisorption results show that 1/2 of the Co atoms are surface atoms, and that the Co clusters are probably affected by the Mn in an electronic fashion.
ST hydrogenation catalyst cobalt silica;
bimetallic cobalt silica catalyst; dispersed catalyst solvated metal prepn; manganese cobalt silica catalyst prepn; iron cobalt silica catalyst prepn
IT Catalysts and Catalysis
Hydrogenation catalysts
Hydrogenolysis catalysts
Isomerization catalysts
(cobalt bimetallic, on silica support, solvated metal atom disperse method in prepn. of)
IT 7439-89-6P, uses and miscellaneous 7439-96-5P, uses and miscellaneous
RL: CAT (Catalyst use); PREP (Preparation); USES (Uses)
(catalyst, with cobalt on silica support, solvated metal atom dispersion method in prepn. of)
IT 7440-48-4, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalysts, prepn. of bimetallic, solvated metal atom dispersion method in)

L26 ANSWER 44 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1985:529759 HCAPLUS
DN 103:129759
TI Effect of alkali metal ions on the catalytic activity of palladium /silica-aluminum phosphate systems

AU Aramendia, M. A.; Borau, V.; Jimenez, C.; Marinas, J. M.; Sempere, M. E.
CS Fac. Cienc., Univ. Cordoba, Cordoba, Spain
SO React. Kinet. Catal. Lett. (1985), 27(1), 133-8
CODEN: RKCLAU; ISSN: 0304-4122
DT Journal
LA English
CC 67-2 (Catalysis, Reaction Kinetics, and Inorganic Reaction Mechanisms)
AB The effects of different alkali metal hydroxides on the **metallic dispersity** and **catalytic activity** of SiO₂-AlPO₄ supported Pd **catalysts** were studied. The redn. by H transfer to benzylideneacetone and styrene (from cyclohexene and n-octanol, resp., as H donors) and the liq.-phase redn. of cyclohexene were used as test reactions to measure the catalytic activity.
ST palladium **catalyst** alkali metal hydroxide; dispersity
palladium oxide **catalyst**; redn **catalyst**
palladium alkali metal; benzylidene acetone redn palladium
catalyst; styrene redn octanol palladium
catalyst; hydrogen transfer catalyst
palladium
IT Hydrogen transfer **catalysts**
Reduction **catalysts**
(palladium-aluminum phosphate-silica, effect of alkali metal ions on activity of)
IT Alkalies
RL: USES (Uses)
(promoters, for palladium-aluminum phosphate-silica **catalysts**)
IT 7440-05-3, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalysts from aluminum phosphate, silica and, effect of alkali metal ions on **dispersity** and **activity** of)
IT 7784-30-7
RL: CAT (Catalyst use); USES (Uses)
(catalysts from palladium, silica and, for redn., effect of alkali metal hydroxides on)
IT 111-87-5, uses and miscellaneous
RL: USES (Uses)
(hydrogen transfer from, to styrene, palladium **catalysts** for)
IT 122-57-6
RL: RCT (Reactant)
(hydrogen-transfer redn. of, by cyclohexene, palladium **catalysts** for)
IT 100-42-5, reactions
RL: RCT (Reactant)
(hydrogen-transfer redn. of, by octanol, palladium **catalysts** for)
IT 110-83-8, reactions
RL: RCT (Reactant)
(redn. of, palladium **catalysts** for)
L26 ANSWER 45 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 1984-133903 [21] WPIX
DNC C1984-056616
TI Dehydrogenating paraffin(s) and naphthene(s) - and dehydro cyclising paraffin(s) using **catalyst** of active carbon contg. uniform dispersion of transition metal (oxide).
DC E19 H04

IN WENNERBERG, A N
 PA (STAD) STANDARD OIL CO INDIANA
 CYC 1
 PI US 4447665 A 19840508 (198421)* 11p
 ADT US 4447665 A US 1983-480407 19830330
 PRAI US 1983-480407 19830330
 IC B01J021-18; C07C005-32
 AB US 4447665 A UPAB: 19930925
 Dehydrogenation of paraffinic and naphthenic cpds. and dehydrocyclisation of paraffins having a straight chain of at least 5 C atoms is carried out at 465-650 (pref. 450-535) deg.C. 2-30 (10-20) atm. and a total WHSV of 0.2-30 (2-10), opt. in the presence of 2-30 (5-15) moles/mole feed of H₂, using a **catalyst** (I) comprising active C having a cage-like structure, a BET surface area of at least 800 (1600, esp. at least 2000) sq.m/g. and a bulk density of at least 0.1 (0.2) g/sq.cm.
 More specifically, (I) contains 0.01-30 wt.% metal and/or metal oxide uniformly dispersed in the porous C matrix, the metal being a transition metal selected from Pt, Pd, Rh, Mo, Cr, W, Fe, Co, Ni and Cu, and the **catalyst** is formed by pyrolysing a uniform powdered mixt. of a precursor (II) of the metal (oxide), a C precursor (III) and inorganic solids (comprising an alkali metal hydroxide) at 400-950 deg.C. in an inert atmosphere.
 (I) exhibits high selectivities and minimised deactivation.
 0/0
 FS CPI
 FA AB
 MC CPI: E10-C04A; E10-C04H; E35-P; E35-Q; E35-U; E35-V; E35-W; E35-X; H04-E03; H04-F02E; J04-F02E; N02; N03-C; N03-D

L26 ANSWER 46 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
 AN 1984-090872 [15] WPIX
 DNC C1984-038504
 TI Prepn. of zeolite type **catalysts** contg. highly dispersed metals - by hydrothermal crystallisation and treatment with hydrogen.
 DC E19 H04 J04
 IN MATHE, T; PETRO, J; TUNGLER, A
 PA (MAGY) MAGYAR TUDOMANYOS AKAD
 CYC 1
 PI HU 29904 T 19840228 (198415)*
 ADT HU 29904 T HU 1981-1699 19810608
 PRAI HU 1981-1699 19810608
 IC B01J023-00; B01J029-04
 AB HU 29904 T UPAB: 19930925
 Prepn. of zeolite-type **catalysts** contg. highly dispersed, catalytically active metals is claimed. The process involves (a) mixing a soln. or dispersion of (i) a cpd. of the formula (R₁(R₃)N(R₂)R₄)⁺X⁻. where R₁₋₄ are ethyl, propyl, butyl, benzyl or hydroxethyl; and X is OH or an inorganic acid residue), (ii) silica, (iii) an aq. soln. contg. Pd, Pt, Ru, Rh, Ag, Cu, Cd, Ni, Fe and/or Re cpds., and opt. (iv) aluminium hydroxide or sodium aluminate; (b) hydrothermally crystallising at 100-200 deg. C for 24-150 hrs; (c) cooling; (d) sepg. the solid prod., and (e) treating the prod. with H₂ at 25-450 deg. C.
 FS CPI
 FA AB
 MC CPI: E05-T; E31-P02; H04-F02; J04-E04; N06-B; N06-E

L26 ANSWER 47 OF 52 HCPLUS COPYRIGHT 2002 ACS
 AN 1985:45434 HCPLUS
 DN 102:45434

TI Effect of metal dispersion on the activity and selectivity of a rhodium/silicon dioxide catalyst for high pressure carbon monoxide hydrogenation
AU Arakawa, Hironori; Takeuchi, Kazuhiko; Matsuzaki, Takehiko; Sugi, Yoshihiro
CS Natl. Chem. Lab. Ind., Yatabe, 305, Japan
SO Chem. Lett. (1984), (9), 1607-10
CODEN: CMLTAG; ISSN: 0366-7022
DT Journal
LA English
CC 23-1 (Aliphatic Compounds)
AB Metal dispersion has a striking influence on the activity and selectivity of Rh/SiO₂ for high pressure CO hydrogenation. With decreasing dispersion, turnover frequency of CO conversion does not change significantly with dispersion up to 0.5, but increases abruptly at a dispersion of 0.45. MeOH is produced selectively at high dispersion. C₂-oxygenated compds. are favored for dispersions of 0.25-0.45. CH₄ formation increases monotonously with a decrease of dispersion.
ST rhodium hydrogenation catalyst; carbon monoxide hydrogenation catalyst; metal dispersion hydrogenation catalyst
IT Hydrogenation catalysts
(rhodium, for carbon monoxide, effect of metal dispersion on)
IT Catalysts and Catalysis
(rhodium, for reaction of carbon monoxide with hydrogen, effect of metal dispersion on)
IT 7440-16-6, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalyst, for high pressure carbon monoxide hydrogenation, effect of metal dispersion on)
IT 630-08-0, reactions
RL: RCT (Reactant)
(high pressure hydrogenation of, over rhodium catalyst, effect of metal dispersion on)
IT 64-17-5P, preparation 64-19-7P, preparation 67-56-1P, preparation
74-82-8P, preparation 75-07-0P, preparation
RL: SPN (Synthetic preparation); PREP (Preparation)
(prepn. of, by hydrogenation of carbon monoxide over rhodium, effect of metal dispersion on)

L26 ANSWER 48 OF 52 HCPLUS COPYRIGHT 2002 ACS
AN 1983:34062 HCPLUS
DN 98:34062
TI Preparation of palladium/aluminum phosphate, palladium /aluminum phosphate-silica, and palladium/aluminum phosphate-.gamma.-alumina and study of their catalytic activity for the reduction of nitrobenzene by hydrogen transfer
AU Aramendia, M. A.; Borau, V.; Jimenez, C.; Marinas, J. M.; Pajares, J. A.
CS Dep. Quim. Org., Univ. Cordoba, Cordoba, Spain
SO J. Catal. (1982), 78(1), 188-96
CODEN: JCTLA5; ISSN: 0021-9517
DT Journal
LA English
CC 22-7 (Physical Organic Chemistry)
Section cross-reference(s): 67
AB The prepn. and catalytic activity for the H-transfer redn. of PhNO₂ of new metallic systems obtained by supporting Pd on AlPO₄, AlPO₄-SiO₂ and AlPO₄-.gamma.-Al₂O₃ are reported. Metal dispersion decreases in the order Pd/AlPO₄-SiO₂ > Pd/AlPO₄ > Pd/AlPO₄-.gamma.-Al₂O₃, and, according to the method of prepn., cationic

exchange > impregnation > anionic exchange. **Catalysts** prepd. by cationic exchange on AlPO₄-SiO₂ show high dispersion values. The **catalysts** have different activities for PhNO₂ redn. to PhNH₂ in refluxing PhNO₂-cyclohexene (383 K), with a selectivity near 100% and without poisoning. No influence of the acidity of the different metal systems was found, the activity depending only on the **dispersion** of the active metal phase.

- ST palladium aluminum phosphate redn **catalyst**; nitrobenzene hydrogen transfer **catalyst**; aniline prepn **catalyst**
- IT Hydrogen transfer
(from cyclohexene to nitrobenzene, effect of **palladium catalysts** on)
- IT Reduction
(of nitrobenzene over **palladium-aluminum phosphate**, mechanism of)
- IT Reduction **catalysts**
(**palladium-aluminum phosphate**, for nitrobenzene, effects of support and method of prepn. on)
- IT 7440-05-3, uses and miscellaneous
RL: PRP (Properties)
(**catalysts** with aluminum phosphate, for nitrobenzene redn., effect of support and prep. method on)
- IT 7784-30-7
RL: CAT (**Catalyst use**); USES (Uses)
(**catalysts**, with **palladium**, for nitrobenzene redn., effect of support and method of prep. on)
- IT 110-83-8, reactions
RL: RCT (Reactant)
(hydrogen transfer from, to nitrobenzene, mechanism of catalytic)
- IT 98-95-3, reactions
RL: RCT (Reactant)
(redn. of, by **hydrogen** transfer from cyclohexene, effect of **palladium-aluminum phosphate catalysts** on)
- L26 ANSWER 49 OF 52 HCAPLUS COPYRIGHT 2002 ACS
AN 1978:169330 HCAPLUS
DN 88:169330
- TI Hydrogenolysis of cyclopentane and **hydrogenation** of benzene on **palladium catalysts** of widely varying dispersion
AU Fuentes, Sergio; Figueras, Francois
CS Inst. Rech. Catal., CNRS, Villeurbanne, Fr.
SO J. Chem. Soc., Faraday Trans. 1 (1978), 74(1), 174-81
CODEN: JCFTAR; ISSN: 0300-9599
- DT Journal
LA English
CC 22-5 (Physical Organic Chemistry)
- AB Adsorption of H and O, titrn. of preadsorbed O, C₆H₆ **hydrogenation**, and cyclopentane hydrogenolysis were studied on **Pd catalysts** of widely varying dispersion. The stoichiometry of O adsorption changes with increasing metallic dispersion. On clean supports free of S and **Fe**, the turnover no. for both hydrocarbon reactions is const. and independent of the **Pd** dispersion. After a suitable redn., S may preferentially inhibit hydrogenolysis, whereas **Fe** preferentially inhibits **hydrogenation**. On an industrial SiO₂ support, the turnover for hydrogenolysis changes with dispersion; this was attributed to contamination of **Pd** by **Fe** from the support.
- ST **palladium catalyst hydrogenation benzene**;

hydrogenolysis cyclopentane catalyst palladium;
adsorption hydrogen oxygen palladium; iron
poison palladium catalyst; sulfur poison
palladium catalyst

IT Adsorption
(of hydrogen and oxygen, on palladium, metallic
dispersion in relation to)

IT Hydrogenation catalysts
(palladium dispersions, for benzene)

IT Hydrogenolysis catalysts
(palladium dispersions, for cyclopentane)

IT 1333-74-0, properties 7782-44-7, properties
RL: PEP (Physical, engineering or chemical process); PROC (Process)
(adsorption of, by palladium, effect of metallic dispersion
on)

IT 7440-05-3, uses and miscellaneous
RL: CAT (Catalyst use); USES (Uses)
(catalysts, for benzene hydrogenation and
cyclopentane hydrogenolysis, effect of metallic
dispersion on activity of)

IT 71-43-2, reactions
RL: RCT (Reactant)
(hydrogenation of, catalyzed by palladium, effect
of metallic dispersion on)

IT 287-92-3
RL: RCT (Reactant)
(hydrogenolysis of, catalyzed by palladium, effect of
metallic dispersion on)

IT 7439-89-6, uses and miscellaneous
RL: USES (Uses)
(poison, for palladium catalysts for
hydrogenation and hydrogenolysis)

IT 7704-34-9, uses and miscellaneous
RL: USES (Uses)
(poison, for palladium catalysts for hydrogenolysis
of cyclopentane)

L26 ANSWER 50 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 1977-11027Y [06] WPIX

TI Methanation of hydrogen and carbon oxide mixts. - using a molten
catalyst resistant to sulphur contg. active
metal catalyst dispersed in zinc or cadmium
halide.

DC E36 H09

PA (SHEL) SHELL OIL CO

CYC 1

PI US 4006177 A 19770201 (197706)*

PRAI US 1975-619543 19751003

IC C07C027-06

AB US 4006177 A UPAB: 19930901
Continuous prodn. of CH₄ from a S-contg. gas mixt. contg. H₂,
CO and/or CO₂ comprises reacting the mixt. at >500 degrees C with
a molten catalyst system consisting of >=1 Zn or Cd halide as
carrier of m.pt. <1000 degrees C in which is dispersed >=1 active
catalyst i.e. finely-divided Fe, Mo, Mn, Ni, Co
, Zn, Ti, Ag, Cu and Th or their oxides and/or carbides. The gaseous
prod. is recovered while some of the carrier is converted to sulphide.
At least a part of the melt is separated (the rest is recycled) and
treated with hydrogen halide at 50-400 degrees C to convert the
sulphide back to halide. The regenerated melt is then recycled.

Method is for mfr. of fuel gas esp. in conjunction with steam gasification of coal. The catalyst is resistant to S so prior desulphurisation is not necessary; less susceptible to C deposition than solid catalysts; and absorbs H₂O so shifts the equilibrium towards CH₄ formation.

FS CPI
FA AB
MC CPI: E10-J02D; H09-C; H09-D; N02; N03; N04-A

L26 ANSWER 51 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 1976-59863X [32] WPIX
TI Reactivation of platinum and palladium catalysts - obtd by treatment with chlorine-air mixts contg hydrogen chloride.
DC H04 J04
PA (KOGL-I) KOGLER H
CYC 1
PI DD 120589 A 19760620 (197632)*
PRAI DD 1968-189217 19681018
IC B01J011-16; B01J023-96
AB DD 120589 A UPAB: 19930901
Activation or re-activation of catalysts contg. Pt or Pd on an alumina or Al silicate carrier is carried out by treatment with a chlorine/air mixt. contg. dry HCl gas in ratio HCl/C₁₂ of 1:3-6, at 300-350 degrees C for 30-70 mins. Used for catalysts used in reforming, isomerising or hydrocracking processes esp. after use and de-activation by catalyst poisons e.g. heavy metals
. Dispersion of the active metal is improved by formation of soluble complexes e.g. H₂PtCl₆. Heavy metals e.g. Pb and As are removed as chlorides in the gas phase. Chloride content after activation is kept within permissible limit of 1% max. Halogenation of the carrier, with loss of activity, as occurs at usual activation temp. of over 450 degrees C, is prevented.

FS CPI
FA AB
MC CPI: H04-F; J04-A02

L26 ANSWER 52 OF 52 WPIX COPYRIGHT 2002 DERWENT INFORMATION LTD
AN 1970-12738R [08] WPIX
TI Low sodium content zeolite based hydroforming - catalysts.
DC H04 J04
PA (SHEL) SHELL INT RES MIJ NV
CYC 6
PI BE 737585 A (197008)*
DE 1941947 A (197009)
NL 6912516 A (197009)
FR 2015902 A (197035)
CA 914138 A (197247)
JP 48022590 B (197328)
DE 1941947 B 19771208 (197750)
NL 161994 B 19791115 (197948)
PRAI GB 1968-39512 19680819
IC B01J029-36; B01J037-02; C10G013-02
AB BE 737585 A UPAB: 19930831
Catalyst for hydroforming petroleum hydrocarbons is based on a zeolite having an SiO₂/Al₂O₃ ratio between 3 and 7 and an alkali metal content reduced to <2, preferably 0.5-0.04% wt. (as NaO). The zeolite is treated with ammonium separately or simultaneously with deposition of a solution of W and/or Mo and a metal from the iron group (VIII) under conditions of pH >6.8, preferably 10-13. Preferably the Na content

of the natural or synthetic zeolite is reduced by precalcination at 250-450 degrees C followed by treatment with boiling 2M NH₄NO₃ solution and calcination at 500-825, preferably 500-700 degrees C. Preferably tungstic acid, ammonium molybdate and a Ni or Co salt are deposited on the zeolite from a basic organic solvent (an alkylamine, alkanolamine, hydrazine or hydroxylamine) and the zeolite is subsequently recalcined at 400-650 degrees C.

Used especially for single stage hydroforming of petroleum hydrocarbons boiling point 350 degrees C by treatment at 250-450 degrees C with 30-200 kg/cm² partial pressure of H₂ at throughputs of 0.2-10 litres hydrocarbon per litre catalyst per hour and 250-5000 litres H₂ per kg hydrocarbon. Also suitable for hydrocracking hydrocarbon liquids including gas oils and asphalt residues. The reduced Na content and good dispersion of the active metals (W, Mo) present results in high catalyst activity, allowing higher than usual feedstock throughputs or use of lower temperatures. The catalysts remain active in the presence of nitrogenous bases and will tolerate some polycyclic aromatic materials in the feedstock, and are easily regenerated by combustion and recalcination if necessary.

FS CPI

FA AB

MC CPI: H04-B03; H04-C03